NEET Exam – 2026 (CODE-12)

Max. Marks: 720 **Duration:** 3 Hours

1. **This paper consists of 180 questions with 3 parts of Physics, Chemistry and Biology**

   - **Physics:** *(Q. No. 1 to 45) Multiple Choice Questions with one correct answer. A correct answer carries 4 Marks. A wrong answer carries a penalty of 1 mark.*
   - **Chemistry:** *(Q. No. 46 to 90) Multiple Choice Questions with one correct answer. A correct answer carries 4 Marks. A wrong answer carries a penalty of 1 mark.*
   - **Biology:** *(Q. No. 91 to 180) Multiple Choice Questions with one correct answer. A correct answer carries 4 Marks. A wrong answer carries a penalty of 1 mark.*

2. **The OMR sheet NEET-2025-180Q is to be used**

3. **Use of calculators and log tables is prohibited**

4. **Darken the appropriate bubble using a pen in the OMR sheet provided to you. Once entered, the answer cannot be changed. Any corrections or modifications will automatically draw a penalty of 1 mark**

5. **No clarification will be entertained during the examination. Doubts in the paper can be reported to the coordinator after the exam**

6. **If the details in the OMR Sheet are not filled, If the OMR sheet is mutilated, torn, white Ink used, the circles filled and scratched, then the OMR sheet will not be graded**

All the best!!

Physics

Multiple Choice Questions with one correct answer. A correct answer carries 4 marks. A wrong answer carries a penalty of 1 mark. 45 x 4 = 180

1. A 100-turn closely wound circular coil of radius 5 cm has a magnetic field of 3.14×10^{-3} T at its centre. The current flowing through the coil, and the magnitude of the magnetic moment of this coil are, respectively: (Take $\mu_0 = 4\pi\times10^{-7}$ T m/A)

   (1) 2.5 A, 2 Am^2 (2) 2.5 A, 20 Am^2 (3) 2 A, 4 Am^2 (4) 2 A, 10 Am^2

   Sol: **Field at the Centre**

   $$B = \frac{\mu_0 ni}{2r}$$

   $$3.14\times10^{-3} = \frac{4\pi\times10^{-7}\times100\times i}{2\times5\times10^{-2}}$$

   $$i = \frac{10\times10^{-5}}{4\times10^{-5}} = 2.5 \text{ A}$$

   Mag moment $M = n\times i\times\pi r^2$

   $M = 100\times2.5\times\pi\times(5\times10^{-2})^2$ A-m^2

   $= 2.5\times3.14\times25\times10^{-4}\times100$

   $= 2.5\times3.14\times25\times10^{-2}$

   $= 1.9625 \approx 2$ A-m^2

   Ans: (1)

2. Match List I with List II:

	List I		List II
A	$E = h\upsilon$	I.	de Broglie wavelength
B	Diffraction and Interference	II.	Particle nature of light
C	$\lambda = h/p$	III.	Wave nature of light
D	Compton effect	IV.	Energy of photon

   Choose the correct answer from the options given below:

   (1) A-IV, B-III, C-I, D-II

   (2) A-I, B – IV, C-III, D-II

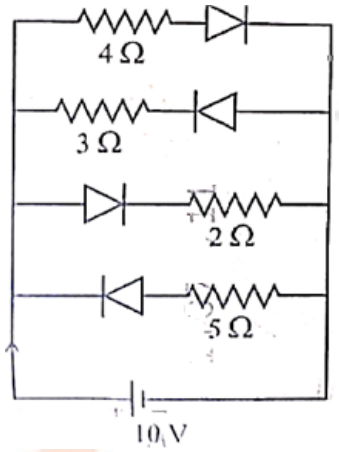
   (3) A-IV, B-I, C-II, D-III

   (4) A-IV, B-III, C-II, D-I

   Sol: Conceptual

   Ans: (1)

3. The current I in the circuit shown below is: (All diodes are ideal dn identical)



(1) $\frac{5}{3}$A (2) $\frac{15}{2}$A (3) $\frac{1}{3}$A (4) $\frac{5}{9}$A

Sol: $i=\frac{10}{\frac{4\times 2}{4+2}}=\frac{10\times 6}{8}=\frac{15}{2}$ A

Ans: (2)

4. The speed of light in vacuum is taken as unity. If light takes 6 min in 40s to reach the earth from the Sun, the distance between the Sun and the Earth in new unit is:

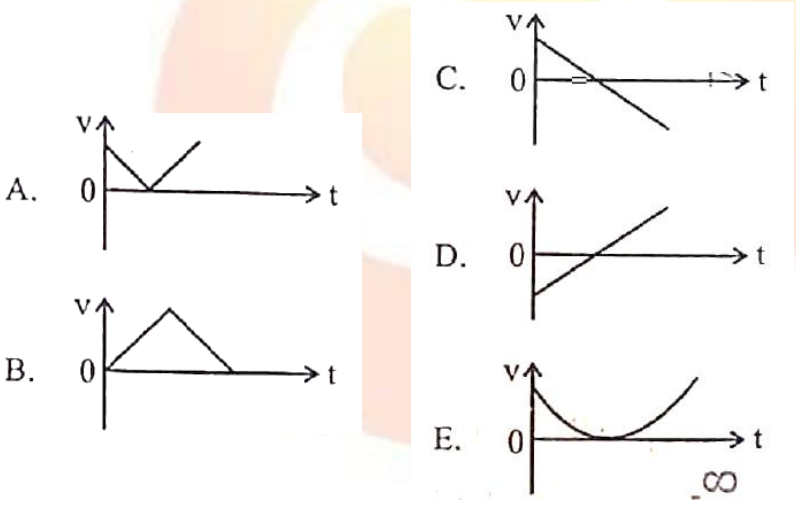
(1) 3×10^{8} (2) 3×10^{10} (3) 400 (4) 500

Sol: $c=1$ unit

distance from the Sun to Earth $=c\times t=1\times 400=400$

Ans: (3)

5. The following plots shown variation of velocity (v) with time (t), of a ball thrown vertically upward, and falling back. Which of the following plots is/are correct?



(1) C only (2) A and E only (3) D only (4) B only

Sol: Conceptual

Ans: (1)

6. In a vernier callipers, 20 VSD coincide with 16 MSD (each division of length 1 mm). The least count of the vernier callipers is:

(1) 0.01 cm (2) 0.1 cm (3) 0.02 cm (4) 0.2 cm

Sol: MSD $= 1\,\text{mm}$

$20\ \text{VSD} = 16\,\text{MSD}$

$$\text{VSD} = \frac{16}{20}\text{MSD}$$

$$\text{LC} = MSD - VSD = MSD - \frac{16}{20}MSD = \frac{4MSD}{20} = \frac{MSD}{5} = \frac{1}{5} = 0.2\,mm = 0.02\,\text{cm}$$

Ans: (3)

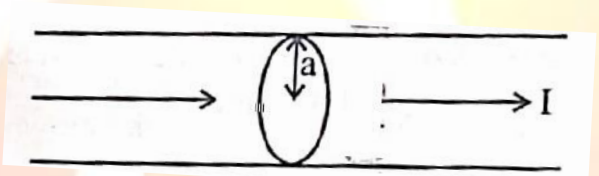
7. An ac circuit contains a resistance of $1\text{k}\Omega$, a capacitor of $0.1\mu\text{F}$ and an inductor of 1mH connected in series. The resonance frequency of the circuit is approximately:

(1) 10.1 kHz (2) 20.7 kHz (3) 15.9 kHz (4) 13.5 kHz

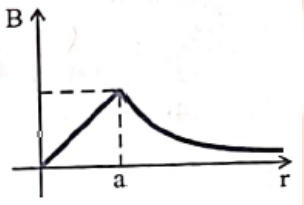
Sol: $f = \dfrac{1}{2\pi\sqrt{LC}} = \dfrac{1}{2\pi\sqrt{1\times10^{-7}\times10^{-3}}} = \dfrac{1}{2\pi\times10^{-5}} = 15.9\ \text{kHz}$

Ans: (3)

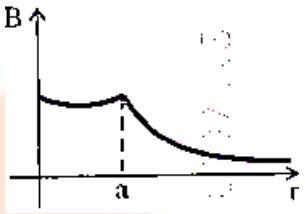
8. The figure given below shows a long straight solid wire of circular cross-section of radius 'a' carrying steady current I. The current I is uniformly distributed across its cross-section. The plot which correctly represents the variation of magnetic field (B) with distance (r) from the axis of the conductor in the region is:



(1)



(2)



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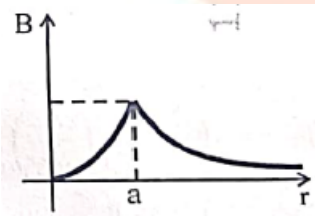
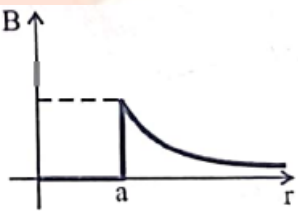


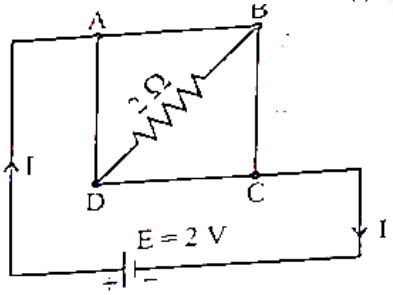
(4)



Sol: Conceptual

Ans: (1)

9. A uniform metallic wire having resistance 4Ω is bent to form a square loop (ABCD) (see figure). A resistance of 2Ω is connected between points B and D and a battery of 2V is connected across points A and C as shown in the figure. Now the value of current (I) is:



(1) 2A (2) 4A (3) 8A (4) 4.5A

Sol: It is a balanced Wheatstone bridge $\therefore I = 2\text{A}$

Ans: (1)

10. An unknown nucleus has a nuclear density of 2.29×10^{17} kg/m^3 and mass of 19.625×10^{-27} kg. Its mass number A is approximately: (take $R_0 = 1.2\times10^{-15}$ m, 4π=12.56)

(1) 12 (2) 19 (3) 20 (4) 16

Sol: $\rho = \dfrac{m}{v}$

$$\rho = \frac{m}{\frac{4}{3}\pi R^3} = \frac{m}{\frac{4}{3}\pi R_0^3 A}$$

$$A = \frac{m}{\frac{4}{3}\pi R_0^3 \rho}$$

$$= \frac{19.926\times10^{-27}}{\frac{4}{3}\times\frac{22}{7}\times\left(1.2\times10^{-15}\right)^3\times2.29\times10^{17}} = 12$$

Ans: (1)

11. A rectangular wire loop of sides 8cm and 3cm with a small cut, is moving out of a region of uniform magnetic field of magnitude 0.3T directed normal to the plane of the loop. The emf developed across the cut, if the velocity of the loop is $2\ \text{cms}^{-1}$, in a direction normal to the shorter side of the loop, will be

(1) 1.8×10^{-4} volt (2) 1.2×10^{-4} volt (3) 1.3×10^{-4} volt (4) 4.8×10^{-4} volt

Sol: $e = Bvl$

$$= \left(\frac{3}{10}T\right)\times\left(2\times10^{-2}\ \text{m/s}\right)\times\left(3\times10^{-2}\ \text{m}\right)$$

$= 1.8 \times 10^{-4}$ v

Ans: (1)

12. A galvanometer of resistance 10Ω gives full scale deflection for a current 1 mA. It is converted into an ammeter of range $0-10$A. The shunt required is:

(1) 0.01Ω 0 (2) 0.10Ω (3) 0.001Ω (4) 1.0Ω

Sol: $S = \dfrac{G}{n-1} = \dfrac{100\Omega}{\dfrac{10}{10^{-3}} - 1} \approx 0.01\Omega$

Ans: (1)

13. In Young's double slit experiment, using monochromatic light of wavelength λ, the intensity of light at a point on the screen where the path difference is λ, is K units. The intensity of light a a point where the path difference is $\dfrac{\lambda}{3}$ will be:

(1) $\dfrac{K}{4}$ (2) K (3) $\dfrac{K}{2}$ (4) 2K

Sol: $\phi_1 = \dfrac{2\pi}{\lambda}\lambda = 2\pi$; $K = 4I_0 \cos^2 \dfrac{\phi_1}{2} = 4I_0 \cos^2 \pi = 4I_0$

$\phi_2 = \dfrac{2\pi}{\lambda}\left(\dfrac{\lambda}{3}\right) = \dfrac{2\pi}{3}$

$K' = 4I_0 \cos^2\left(\dfrac{\phi_2}{2}\right) = 4I_0 \cos^2\left(\dfrac{\pi}{3}\right)$

$= \dfrac{4I_0}{4} = \dfrac{K}{4}$

Ans: (1)

14. The magnitude and direction of the acceleration produced in a body of mass 5kg when two mutually perpendicular across 8N and 6N act on it, are respectively:

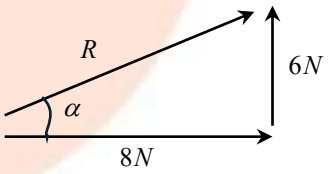
(1) $2\,\text{ms}^{-2}; \tan^{-1}(3/4)$ with 6N force (2) $2\,\text{ms}^{-2}; \tan^{-1}(4/3)$ with 8N force

(3) $2\,\text{ms}^{-2}; \tan^{-1}(3/4)$ with 8N force (4) $20\,\text{ms}^{-2}; \tan^{-1}(4/3)$ with 8N force

Sol:

$R = \sqrt{6^2 + 8^2}$

$= 10$

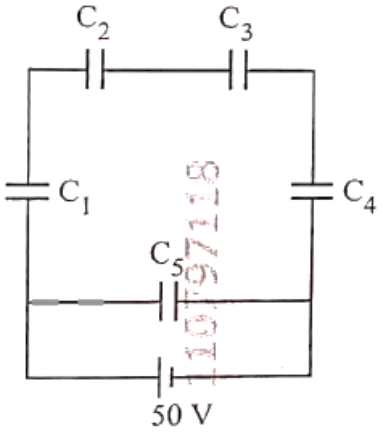


$\tan\alpha = \dfrac{6}{8} = \dfrac{3}{4}$

$\alpha = \tan^{-1}\dfrac{3}{4}$ from 8N

Ans: (3)

15. Five capacitors of capacitances $C_1 = C_2 = C_3 = C_4 = 10\,\mu F$ and $C_5 = 2.5\,\mu F$ are connected as shown, along with a battery of 50V.



The equivalent capacitance and the charges on each capacitor respectively are:

(1) $5\,\mu F, 125\,\mu C$ on all capacitors

(2) $5\,\mu F, 250\,\mu C$ on all capacitors

(3) $4\,\mu F, 250\,\mu C$ on C_1 to C_4 and $125\,\mu C$ on C_5

(4) $5\,\mu F, 125\,\mu C$ on C_1 to C_4 and $25\,\mu C$ on C_5

Sol: $C_e = \frac{C}{4} + C_5$

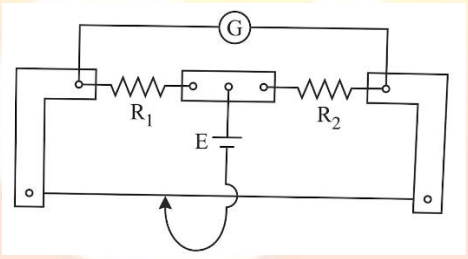
$C_e = \frac{10}{4} + 2.5 = 5\mu F$

$Q = CeV = 250\,\mu C$

Charge in each capacitor $= 125\,\mu C$

Ans: (1)

16. In a metre bridge experiment (see figure), the positions of the cell, E, and galvanometer, G, are interchanged We shall observe in the galvanometer :



(1) Only the right-sided deflection

(2) Only the left-sided deflection

(3) There will be no deflection irrespective of the position of the jockey

(4) Both right-sided and left-sided deflection and at balance point, no deflection.

Sol: Conceptual

$\rightarrow$ If the bridge is initially balanced, it will still remain balance

$\rightarrow$ If we move jockey galvanometer will show deflection in both directions.

Ans: (4)

17. The power of a crane, which lifts a mass of 1000 kg to a height of 20 m in 10 s is : ($g = 9.8$ m/s^2)

(1) 19.6 W (2) 39.2 W (3) 39.2 kW (4) 19.6 kW

Sol: $P = \dfrac{mgh}{\Delta t} = \dfrac{(10^3 \text{ kg}) \times (9.8 \text{ m/s}^2) \times (20 \text{ m})}{10 \text{ s}} = 19.6 \text{ kW}$

Ans: (4)

18. Match List I with List II.

List I	List II
A. Young's Modulus	I. $\dfrac{\Delta d}{\Delta L}\left(\dfrac{L}{d}\right)$
B. Compressibility	II. $\dfrac{FL}{A(\Delta L)}$
C. Bulk Modulus	III. $-\dfrac{1}{\Delta P}\left(\dfrac{\Delta V}{V}\right)$
D. Poisson's Ratio	IV. $-P\left(\dfrac{V}{\Delta V}\right)$

Choose the correct answer from the options given below :

(1) A-I, B-IV, C-III, D-II (2) A-IV, B-I, C-II, D-III

(3) A-III, B-II, C-I, D-IV (4) A-II, B-III, C-IV, D-I

Sol: A. $Y = \dfrac{FL}{A\Delta L} \longrightarrow$ II

B. $C = -\dfrac{1}{\Delta P}\left(\dfrac{\Delta V}{V}\right) \longrightarrow$ III

C. $K = -P\left(\dfrac{V}{\Delta V}\right) \longrightarrow$ IV

D. $\mu = \dfrac{\Delta d}{\Delta L}\left(\dfrac{L}{d}\right) \longrightarrow$ I

Ans: (4)

19. In a concave lens, a ray of light emanating from the object parallel to the principal axis of the lens, after refraction:

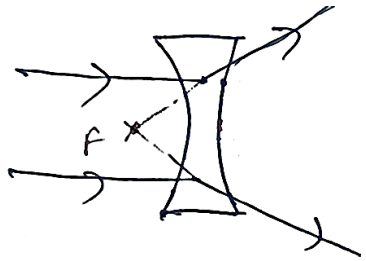
(1) emerges parallel to the principal axis.

(2) appears to diverge from the first principal focus.

(3) passes through 2F, which is the radius of curvature of the lens.

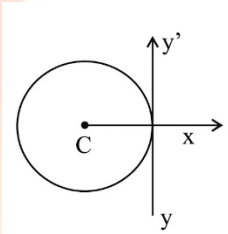
(4) passes through the second principal focus.

Sol:



Ans: (2)

20. A thin wire of length 'L' and linear mass density 'm' is bent into a circular ring (in x-y plane) with centre 'C' as shown in figure. The moment of inertia of the ring about an axis yy' will be :



(1) $\frac{3\,mL^3}{8\,\pi^2}$ (2) $\frac{3\,mL^3}{8\,\pi}$ (3) $\frac{3\,mL^2}{8\,\pi^2}$ (4) $\frac{3\,mL^2}{8\,\pi}$

Sol: $I = I_{cm} + md^2$

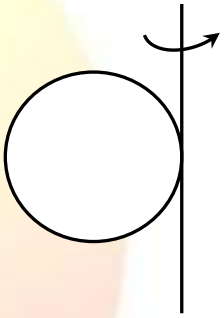
$I = \frac{MR^2}{2} + MR^2$

$= \frac{3MR^2}{2}$ ---(1)

$M = m \times L$ ---(2)

$2\pi R = L$

$R = \frac{L}{2\pi}$ ---(3)



② and ③ in ①

$I = \frac{3}{2} \times (m \times L) \times \frac{L^2}{4\pi^2} = \frac{3mL^3}{8\pi^2}$

Ans: (1)

21. Each side of a metallic cube of mass 5.580 kg is measured to be 9.0 cm. Keeping the significant figures in view, the density of the material of the cube can be best expressed as $X \times 10^3$ kg m^{-3}, where the value of X is :

(1) 7.654 (2) 7.7 (3) 7.65 (4) 7.6

Sol: $V = 729\ \text{cm}^3$

$M = 5.580\ \text{kg}$

$\rho = \frac{M}{V} = \frac{5.580}{729}$

$= 0.007654\ \text{g/cm}^3$

$= 7.654\ \text{kg/m}^3 \approx 7.6\ \text{kg/m}^3$ (According to NCERT)

Ans: (4)

22. For a travelling harmonic wave $y(x,t) = 2.0\cos 2\pi(10\ \text{t} - 0.0080\ \text{x} + 0.35)$, where x and y are in cm and t in s. The phase difference between oscillatory motion of two points separated by a distance of 0.5 m is :

(1) 8π rad (2) 0.08π rad (3) 0.008π rad (4) 0.8π rad

Sol: $\Delta\phi = \frac{2\pi}{\lambda}\Delta x$

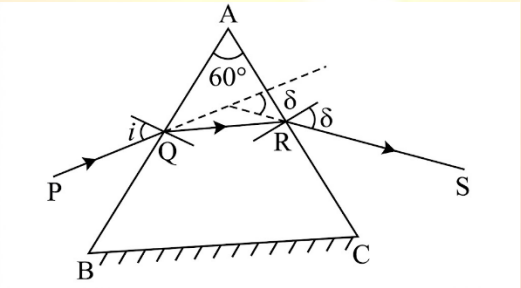
$= k\Delta x$

$= 2\pi \times 0.0080 \times 0.5 \times 100$

$= 0.8\pi$ rad

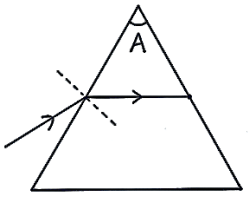
Ans: (4)

23. A ray of monochromatic light is passing through an equilateral prism (ABC) as shown in the figure. The refracted ray (QR) is parallel to its base (BC) and the angle of incidence (i) is 50°. Then the angle of deviation (δ) is :



(1) 40° (2) 45° (3) 55° (4) 35°

Sol: $\Rightarrow$ Since ray inside prism is parallel to base is a minimum deviation case $\delta = i + e - A$



$= 2i - A \quad [i = e]$

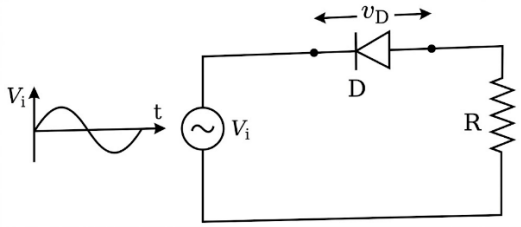
$= 2 \times 50^\circ - 60^\circ$

$= 100^\circ - 60^\circ$

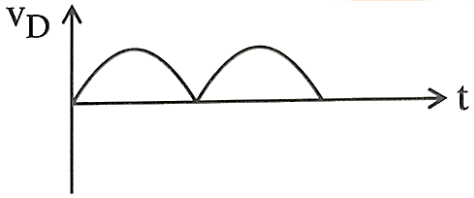
$= 40^\circ$

Ans: (1)

24. In the circuit shown below, the voltage appearing across the diode D will be of the form :



(1)



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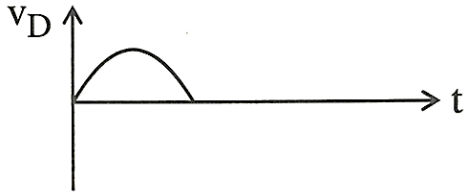

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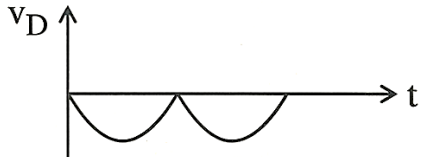

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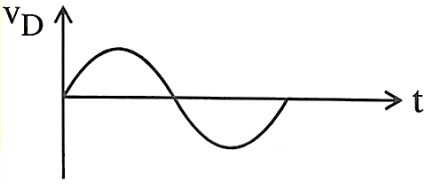
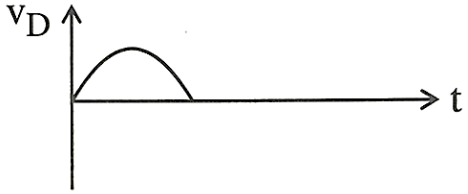


Sol:

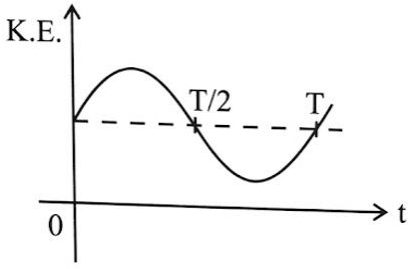


Half wave rectifier.

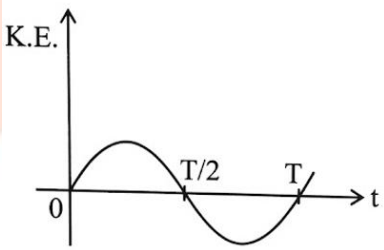
Ans: (2)

25. For a simple pendulum, having time period T, the variation of kinetic energy (K.E.) with time (t) is represented by :

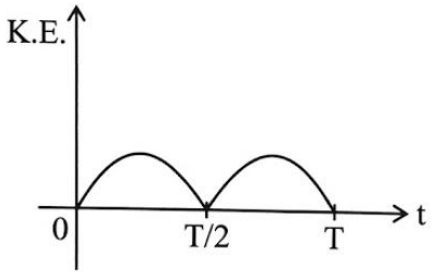
(1)



(2)



(3)



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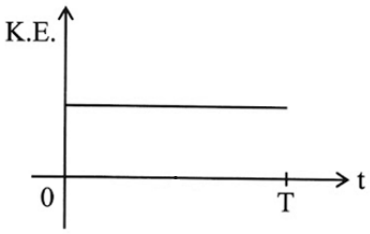


Sol: Conceptual : Time period of K.E = T/2

Ans: (3)

26. A resistor is connected to a battery of $12V$ emf and internal resistance 2Ω. If the current in the circuit is $0.6A$, the terminal voltage of the battery is:

(1) $10V$ (2) $10.8V$ (3) $12V$ (4) $1.2V$

Sol: $V = E - Ir$

$= 12 - 0.6 \times 2$

$= 12 - 1.2$

$= 10.8$ Volt

Ans: (2)

27. The amount of work done to raise a mass 'm' from the surface of the Earth to a height equal to the radius of the Earth 'R', will be:

(1) $2mgR$ (2) mgR (3) $mg\frac{R}{4}$ (4) $mg\frac{R}{2}$

Sol: $W = \Delta U$

$= \frac{-GMm}{2R} - \left(\frac{-GMm}{R}\right)$

$= \frac{GMm}{2R} = \frac{mgR}{2}$

Ans: (4)

28. An electric heater supplies heat to a system at a rate of $100W$. If the system performs work at a rate of $75J/s$, then the rate at which internal energy increases will be:

(1) $125W$ (2) $100W$ (3) $25W$ (4) $75W$

Sol: $\frac{\Delta Q}{\Delta t} = \frac{\Delta U}{\Delta t} + \frac{\Delta W}{\Delta t}$

$100 = -+75 \Rightarrow \frac{\Delta U}{\Delta t} = 25J/s = 25W$

Ans: (3)

29. A room heater is rated $400W, 220V$. If the supply voltage drops to $200V$, what will be the power consumed (approximately)?

(1) $121W$ (2) $331W$ (3) $200W$ (4) $400W$

Sol: $P_{\text{cons}} = \left(\frac{V}{V_{\text{Rat}}}\right)^2 P_{\text{rated}}$

$= \left(\frac{200}{220}\right)^2 \times 400W$

$= \frac{400}{484} \times 400W = 331W$

Ans: (2)

30. When a ruler falls vertically, 5 different persons catch it with different reaction times. $\left(g = 9.8\,\text{ms}^{-2}\right)$

A. Person A has reaction time of $0.20s$.

B. Person B has reaction time of $0.22s$.

C. Person C has reaction time of $0.18s$.

D. Person D has reaction time of $0.19s$.

E. Person E has reaction time of $0.21s$.

What is the correct order of the distance travelled by the ruler for each person?

(1) C > D > A > B > E (2) C > D > A > E > B (3) B > E > A > C > D (4) B > E > A > D > C

Sol: Distance = Speed $\times \Delta t$

B > E > A > D > C

Ans: (4)

31. Consider two uncharged capacitors of equal capacitance $200\,\text{pF}$. One of them is charged by a $100V$ supply and disconnected. Now this capacitor is connected to the uncharged capacitor. The amount of electrostatic energy lost in the process is:

(1) $1.0\times10^{-6}\,J$ (2) $0.5\times10^{-6}\,J$ (3) $0.5\,J$ (4) $1.0\,J$

Sol: $V_{\log t} = \frac{1}{2}\frac{C_1C_2}{C_1+C_2}(V_1 - V_2)^2$

$= \frac{1}{2}\left(\frac{C}{2}\right)V^2$

$= \frac{1}{4}\times200\times10^{-12}\times10^4$

$= 0.5\times10^{-6}J$

Ans: (2)

32. Savitha, a XI standard student, while conducting an experiment to determine the effective length of a simple pendulum L, notes down the data of time taken to complete 30 oscillations as 60s and hence calculates the length of the simple pendulum as: (Take $\pi^2 = 9.8$, and $g = 9.8\,\text{m/s}^2$)

(1) 2 m (2) 0.75 m (3) 1.5 m (4) 1 m

Sol: $T_{osc} = \frac{60\,\text{sec}}{30} = 2\,\text{sec} = 2\pi\sqrt{\frac{l}{g}}$

$l = 1\,m$

Ans: (4)

33. The peak value of an alternating current is $5A$ and frequency is $60\,Hz$. How long will the current, starting from zero, take to reach the peak value?

(1) $\frac{1}{240}$s (2) $\frac{1}{30}$s (3) $\frac{1}{120}$s (4) $\frac{1}{60}$s

Sol: $t = \frac{T}{4} = \frac{1}{4f} = \frac{1}{240}\,\text{sec}$

Ans: (1)

34. In interference and diffraction, the light energy is redistributed. If it reduces in one region, producing a dark fringe, it increases in another region, producing a bright fringe.

A. As there is no gain or loss of energy, these phenomena are consistent with the principle of conservation of energy.

B. Diffraction and interference are characteristics exhibited only by light waves.

Choose the correct answer from the options given below:

(1) A is true, but B is false (2) A is true and B is also true

(3) A is false, but B is true (4) Both A and B are false

Sol: A is true, but B is false

Ans: (1)

35. A box of mass $15\,kg$ is kept on the floor of a stationary trolley. The coefficient of static friction between the box and the trolley is 0.12. Keeping the box in stationary state over the trolley, the maximum acceleration with which the trolley can be moved horizontally in ms^{-2} is: $\left(g = 10\,\text{m/s}^2\right)$

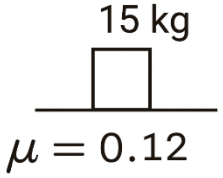
(1) 1.5 (2) 1.8 (3) 2.1 (4) 1.2

Sol: $a_{\max} = \mu g$

$= 0.12 \times 10$

$= 1.2\,m/s^2$

Ans: (4)



36. The sum of kinetic energy and potential energy of a simple pendulum bob is 0.02 *joule*. The speed of the simple pendulum bob at equilibrium position is approximately:

(Consider mass of the bob = 20g)

(1) 1.41 m/s (2) 14.1 m/s (3) 0.2 m/s (4) 2.0 m/s

Sol: $TE = 6.02$

At equilibrium

$TE = KE$

$\frac{1}{2} \times m \times v^2 = 0.02J$

$v = \sqrt{2}\, m/s = 1.41 m/s$

Ans: (1)

37. Four statements are given (A is mass number):

A. The volume of a nucleus is proportional to $A^{1/3}$

B. The volume of a nucleus is proportional to A.

C. The difference in mass of an atom and its nucleus is called the mass defect.

D. The difference in mass of a nucleus and its constituents is called the mass defect.

Choose the correct answer from the options given below :

(1) B and D are true, but A and C are false

(2) A and D are true, but B and C are false

(3) A and C are true, but B and D are false

(4) B and C are true, but A and D are false

Sol: Volume of nucleus $\propto A$

B is true

and D is also true

$\Delta m = \left(ZM_p + (A - Z) M_n - M \right)$

Ans: (1)

38. The angular speed of a flywheel is increased from 600 rpm to 1200 rpm in 10 s. The number of revolutions completed by the flywheel during this time is :

(1) 600 (2) 900 (3) 300 (4) 150

Sol: $W_f = 1200\, rpm$

$= 1200 \times \frac{2\pi}{60} rad/s$

$= 40\pi$

$w_i = 600\ rpm$

$w_i = 20\ rpm$

$$\theta = \frac{\omega_f^2 - \omega_1^2}{2\left(\frac{\omega_f - \omega_i}{t}\right)}$$

$$\theta = \frac{\omega_f + \omega_1}{2} \times t$$

$$= \frac{40\pi + 20\pi}{2} \times 10$$

$$\theta = -300\pi$$

Number of revolutions $= \frac{300\pi}{2\pi} = 150$

Ans: (4)

39. A submarine is designed to withstand an absolute pressure of 100 atm. How deep can it go below the water surface?

(Consider the density of water $= 1000 \text{ kg m}^{-3}$, $1 \text{ atm} = 1 \times 10^5 \text{ Pa}$ and gravitational acceleration $g = 10 \text{ m/s}^2$)

(1) 9900 m (2) 99 m (3) 9000 m (4) 990 m

Sol: $p = p_0 + 8gh$

$$100 \times 10^5 = 10^5 + 10^3 \times 10 \times h$$

$$\frac{49 \times 10^5}{10^4} = h; \; h = 990m$$

Ans: (4)

40. Match List I with List II:

	List I (Electromagnetic wave)		List II (Production)
A.	Microwave	I.	Electrons in atoms emit light when they move from a higher energy level to a lower energy level
B.	Visible light	II.	Radioactive decay of nucleus
C.	Gamma rays	III.	Vibration of atoms and molecules
D.	Infrared rays	IV.	Klystron valve or magnetron valve

Choose the correct answer from the option given below:

(1) A-III, B-I, C-II, D-IV

(2) A-III, B-IV, C-I, D-II

(3) A-IV, B-III, C-II, D-I

(4) A-IV, B-I, C-II, D-III

Sol:

A $\rightarrow$ (IV)

B $\rightarrow$ (I)

C $\rightarrow$ (II)

D $\rightarrow$ (III)

Ans: (4)

41. Which of the following statements are correct ?

A. Inside a conductor, the electrostatic field is zero.

B. Electric field at the surface of a charged conductor does not depend on its surface charge density.

C. The interior of a charged conductor can have no excess charge in the static situation.

D. At the surface of a charged conductor, the electrostatic field must be normal to the surface at every point.

E. The electrostatic potential is zero everywhere inside a charged conductor.

Choose the correct answer from the options given below :

(1) C, D and E only (2) A, B and D only

(3) A, C and D only (4) A, C and E only

Sol: A, C and D only

Ans: (3)

42. For a metal of work function 6.6 eV, which of the following wavelengths of incident radiation does *not* give rise to the photoelectric effect ?

(Take Planck's constant as 6.6×10^{-34} J/ s)

(1) 200 nm (2) 150 nm (3) 100 nm (4) 50 nm

Sol: $E=\dfrac{1240}{\lambda(\text{in } nm)}ev$

$\lambda=200$; $E<\phi<6.6$

Ans: (1)

43. In the first excited state of hydrogen atom, the energy of its electron is -3.4 eV . The radial distance of the electron from the hydrogen nucleus in this case is approximately :

(Take 1 eV $=1.6\times10^{-19}$ J , $e=1.6\times10^{-19}$ C and $\dfrac{1}{4\pi\epsilon_0}=9\times10^9$ N m^2 / C^2)

(1) 2.1×10^{-8} m (2) 2.1×10^{-11} m (3) 2.1×10^{-9} m (4) 2.1×10^{-10} m

Sol: $r_n=0.529\times10^{-10}n^2m$

$n=2$

$r_2=2.1\times10^{-10}m$

Ans: (4)

44. Two statements are given below :

A. When the forward bias voltage across a p-n junction diode increases above a certain threshold voltage, the diode current increases significantly.

B. This current 'is' called reverse saturation current.

Choose the correct answer from the options given below :

(1) Both Statements A and B are false

(2) Statement A is true, but Statement B is false

(3) Both Statements A and B are true

(4) Statement A is false, but Statement B is true

Sol: Statement A is True :- when the forward bias voltage exceeds.

The threshold (knee) voltage, the potential barrier is reduced significantly, allowing a high forward current to flow with increase rapidly

Statement (B) is false: - the current that flow when a p-n junction is forward biased is called forward current.

Ans: (2)

45. A flask contains argon and chlorine in the ratio of 2 : 1 by mass. The temperature of the mixture is 27°C.

The ratio of root mean square speed of the molecules of the two gases $\left(\frac{v_{rms}^{Ar}}{v_{rms}^{Cl}}\right)$ is :

(Atomic mass of argon = 40.0 u and molecular mass of chlorine = 70.0 u)

(1) $\frac{\sqrt{7}}{2}$ (2) $\frac{7}{2}$ (3) $\frac{7}{4}$ (4) $\frac{2}{\sqrt{7}}$

Sol:

$$\frac{V_{Ar}}{V_{ci}} = \sqrt{\frac{Mcl}{MAr}}$$

$$= \sqrt{\frac{70}{40}}$$

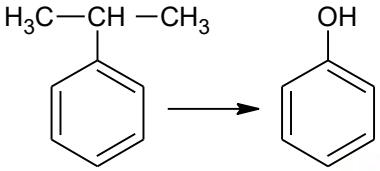
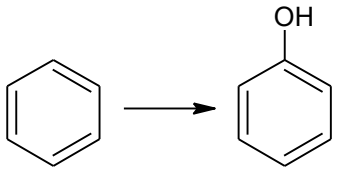
$$= \frac{\sqrt{7}}{2}$$

Ans: (1)

Chemistry

Multiple Choice Questions with one correct answer. A correct answer carries 4 marks. A wrong answer carries a penalty of 1 mark. 45 x 4 = 180

46. Match List I with List II:

	List – I		List – II
A.		I.	(i) oleum; (ii) $NaOH, \Delta$; (iii) H^+
B.	$CH_3COOH \rightarrow CH_3CH_2OH$	II.	(i) O_2 ; (ii) H_2O / H^+
C.	$CH_3CH_2CH_2OH \rightarrow CH_3 - \underset{\underset{OH}{\vert}}{C}H - CH_3$	III.	(i) CH_3OH, H^+ ; (ii) H_2 , catalyst
D.		IV.	(i) conc. H_2SO_4, Δ ; (ii) H^+ / H_2O

Choose the correct answer from the options given below.

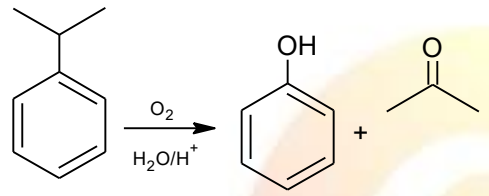
(1) A – I, B – III, C – IV, D – II

(2) A – II, B – IV, C – III, D – I

(3) A – II, B – III, C – I, D – IV

(4) A – II, B – III, C – IV, D – I

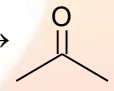
Sol:

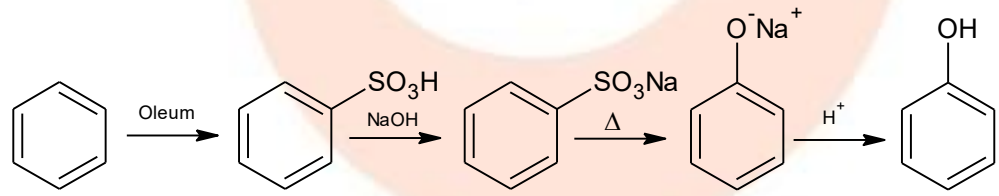


Cumene hydroperoxide synthesis

$$CH_3COOH + CH_3OH \xrightarrow{H^+} CH_3COOCH_3 \xrightarrow[\text{catalyst}]{H_2} CH_3CH_2OH$$

$$CH_3 - CH_2 - CH_2 - OH \xrightarrow[H_2SO_4]{\text{conc.}} CH_3 - CH = CH_2 \xrightarrow[H_2O]{H^+}$$





Ans: (4)

47. The major product Z formed in the following sequence of reactions is

$$C_2H_6 \xrightarrow[\text{UV light}]{Cl_2} \underset{\text{monochlorinated}}{X} \xrightarrow{NH_3} Y \xrightarrow[(ii)\,H_2O]{(i)\,NaNO_2, HCl} Z$$

(1) $C_2H_5 - N = N - OH$ (2) C_2H_5OH (3) $C_2H_5NO_2$ (4) $C_2H_5NH_2$

Sol: $C_2H_6 \xrightarrow[\text{UV light}]{Cl_2} CH_3CH_2Cl \xrightarrow[-HCl]{NH_3} CH_2CH_2NH_2 \xrightarrow[HCl,\ H_2O]{NaNO_2} CH_3CH_2OH$

Ans: (2)

48. In a qualitative analysis, Bi^{3+} is detected by appearance of precipitate of $BiO(OH)(s)$. Calculate pH when the following equilibrium exists at 298 K:

$BiO(OH)(s) \rightleftharpoons BiO^+(aq) + OH^-(aq)$

$K = 4\times10^{-10}$ (given: $\log 2 = 0.3010$)

(1) 4.699 (2) 8.714 (3) 9.301 (4) 5.286

Sol: $Bio(OH) \rightleftharpoons Bio^+ + OH^-$, $K_{sp} = 4\times10^{-10}$, $s \approx 2\times10^{-5}$ in normal solution.

$\left[OH^-\right] = 2\times10^{-5}$

$pOH = -\log\left(2\times10^{-5}\right) = 5 - \log 2 = 4.7$

$pH = 14 - 4.7 = 9.3$

Ans: (3)

49. When 1 dm^3 of CO_2 gas is passed over hot coke, the volume of gaseous mixture after complete reaction at STP becomes 1.4 dm^3. The composition of the gaseous mixture at STP is

(1) $0{\cdot}6\ \text{dm}^3$ of CO, $0{\cdot}8\ \text{dm}^3$ of CO_2 (2) $0{\cdot}8\ \text{dm}^3$ of CO, $0{\cdot}8\ \text{dm}^3$ of CO_2

(3) $0{\cdot}8\ \text{dm}^3$ of CO, $0{\cdot}6\ \text{dm}^3$ of CO_2 (4) $0{\cdot}6\ \text{dm}^3$ of CO, $0{\cdot}4\ \text{dm}^3$ of CO_2

Sol: $CO_2 + C \rightarrow 2CO$

1 vol 2 vol

$1-x$ $2x$

$1 - x + 2x = 1.4$

$1 + x = 1.4 \quad \Rightarrow x = 0.4$

$CO = 2x = 0.8\,dm^3$

$CO_2 = 1 - x$

$= 1 - 0.4 = 0.6\,dm^3$

Ans: (3)

50. Match List I with List II :

	List I (Quantum Numbers)		List II (Orbital)
	'n' '1'		
A	2 1	I.	3d
B	4 0	II.	2p
C	5 3	III.	4s
D	3 2	IV.	5f

Choose the **correct** answer from the options given below:

(1) A-II, B-III, C-IV, D-I (2) A-I, B-II, C-III, D-IV

(3) A-IV, B-II, C-III, D-I (4) A-II, B-III, C-I, D-IV

Sol: $n = 2, l = 1$ $: 2p$

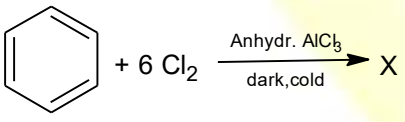
$n = 4, l = 0$ $: 4s$

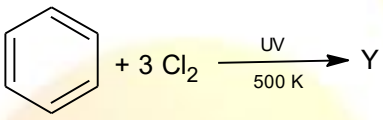
$n = 5, l = 3$ $: 5f$

$n = 3, l = 2$ $: 3d$

Ans: (1)

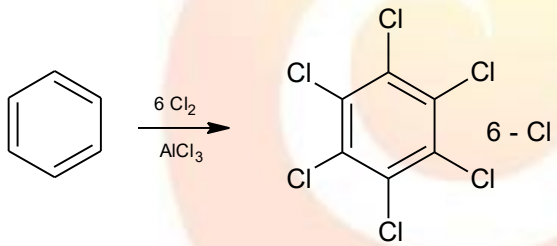
51. The number of chlorine atoms present in the organic products X and Y of the following reactions, respectively, are :

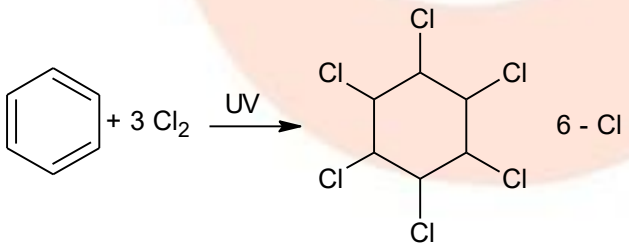




(1) 3 and 6 (2) 6 and 6 (3) 6 and 3 (4) 3 and 3

Sol:





Ans: (2)

52. In the following reaction sequence, X and Z, respectively are :

$$CH_3CH_2CH_2-OH+PCl_5 \rightarrow CH_3CH_2CH_2Cl+X+HCl \xrightarrow{\text{alc.KOH}\,\Delta} Y \xrightarrow[(C_6H_5CO)_2O_2]{HBr} Z$$

(1) $X = POCl_3; Z = CH_3-\underset{\underset{Br}{|}}{CH}-CH_3$

(2) $X = H_3PO_3; Z = CH_3CH_2CH_2-Br$

(3) $X = H_3PO_3; Z = CH_3-\underset{\underset{Br}{|}}{CH}-CH_3$

(4) $X = POCl_3; Z = CH_3CH_2CH_2-Br$

Sol:

$$CH_2-CH_2-CH_2-CH \longrightarrow CH_3-CH_2-CH_2-Cl+POCl_3+HCl$$

$$\xrightarrow{\text{alc.KOH}/\Delta} CH_3-CH=CH_2 \xrightarrow[\text{Peoxide}]{HBr} CH_3-CH_2-CH_2-Br$$

Ans: (4)

53. Match List I with List II :

	List I (Transition metal/compound/complex)		**List II (Catalytic Role)**
A	V_2O_5	I	Preparation of ammonia from N_2/H_2 mixture
B	Fe	II	Polymerisation of alkynes
C	$PdCl_2$	III	Preparation of H_2SO_4 from SO_2
D	Ni complex	IV	Oxidation of ethyne to ethanal

Choose the **correct** answer from the options given below :

(1) A-III, B-IV, C-I, D-II

(2) A-II, B-I, C-IV, D-III

(3) A-IV, B-I, C-III, D-II

(4) A-III, B-I, C-IV, D-II

Sol:

$V_2O_5 \rightarrow$ Preparation of H_2SO_4

Fe $\rightarrow$ Haber's process

$PdCl_2 \rightarrow$ oxidation of ethylene to ethanal

Ni complex $\rightarrow$ polymerisation of alkyne

Ans: (4)

54. Calculate emf of the half cell given below:

$Pt\ (s)|\ H_2\ (g,\ 2\ atm)|\ HCl\ (aq,\ 0{\cdot}02\ M)$

$E^{\circ}_{H_2/H^+} = 0$ V

(Given : $\frac{2{\cdot}303\ RT}{F} = 0{\cdot}059$, $\log 2 = 0{\cdot}3010$)

(1) 0·109 V

(2) 0·035 V

(3) –0·035 V

(4) –0·109 V

Sol: $Pt\ (s)|\ H_2\ (g,\ 2\ atm)|\ HCl\ (aq,\ 0\cdot02\ M)$

$H_{2(g)} \rightarrow 2H^+_{(aq)} + 2e^-$

$$E_{cell} = E^0_{cell} - \frac{0.059}{2}\log\frac{\left[H^+\right]^2}{pH_2} = 0 - 0.0295\log\left[\frac{(0.02)^2}{2}\right] = -0.0295\log\left(\frac{4\times10^{-4}}{2}\right)$$

$$= -0.0295\left(-4 + \log 2\right) = -0.0295\times -3.7 = 0.109V$$

Ans: (1)

55. Identify the correct statement about ClF_3 from the following options :

(1) It has a trigonal pyramidal geometry with two lone pairs on Cl atom.

(2) It has T-shaped geometry with two lone pairs on Cl atom.

(3) It has a planar trigonal geometry with two lone pairs on Cl atom.

(4) It has T-shaped geometry with three lone pairs on Cl atom.

Sol: ClF_3 : sp^3d hybridisation: T-shaped geometry with two lone pairs on Cl atom

Ans: (2)

56. Match List I with List II :

	List I (Order of reaction)		**List II (Unit of rate constant)**
A.	Zero order	I	$mol^{-1}\ L\ s^{-1}$
B.	First order	II	$mol^{-2}\ L^2\ s^{-1}$
C.	Second order	III	s^{-1}
D.	Third order	IV	$mol\ L^{-1}\ s^{-1}$

Choose the **correct** answer from the options given below :

(1) A-IV, B-III, C-II, D-I (2) A-I, B-II, C-III, D-IV

(3) A-IV, B-III, C-I, D-II (4) A-IV, B-II, C-I, D-III

Sol: Zero order - $mol\ L^{-1}\ s^{-1}$

First order – s^{-1}

Second order – $mol^{-1}\ L\ s^{-1}$

Third order – $mol^{-2}\ L^2\ s^{-1}$

Ans: (3)

57. The calculated 'spin-only' magnetic moment of $Ti^{2+}(3d^2)$ is :

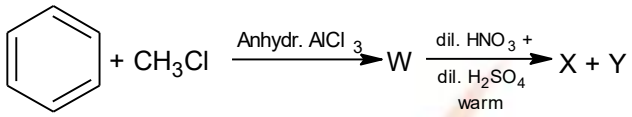
(1) 2.84 BM (2) 5.92 BM (3) 4.90 BM (4) 3.87 BM

Sol: $Ti^{2+}:[Ar]3d^2 4s^0$

Magnetic moment: $\sqrt{n(n+2)} = \sqrt{8} = 2.84$

Ans: (1)

58. Two products X and Y are formed in the following reaction sequence.

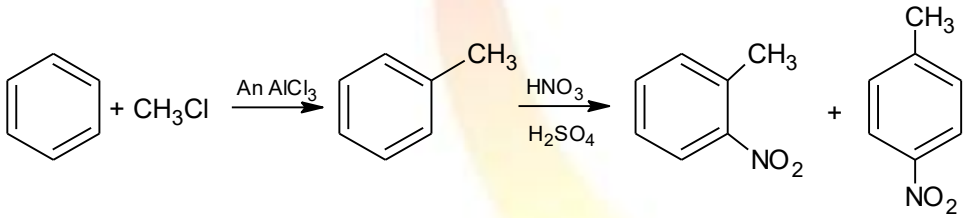


The suitable method that can be used for the separation of products X and Y is :

(1) Continuous extraction (2) Differential extraction

(3) Fractional distillation (4) Sublimation

Sol:



Separated by Fractional distillation

Ans: (3)

59. A bulb is rated at 150 watt, converting 8% energy into light. If energy of one photon is $4{\cdot}42\times10^{-19}$ J, how many photons are emitted by the bulb per second?

(1) $1{\cdot}35\times10^{19}$ (2) $4{\cdot}06\times10^{19}$ (3) $2{\cdot}71\times10^{19}$ (4) $27{\cdot}2\times10^{19}$

Sol: 150 watt = 150 J

$E = nh\nu$

Energy emitted / sec = $\frac{150\times8}{100} = 12 \text{ J s}^{-1}$

Number of photons = $\frac{12}{4.42\times10^{-19}} = 2.71\times10^{19}$ s

Ans: (3)

60. In a test tube containing a salt, a few drops of dilute H_2SO_4 was added, which gave colourless vapours having the smell of vinegar. The vapours turned the blue litmus paper red.

Identify the **correct** anion from the following :

(1) Acetate, CH_3COO^- (2) Carbonate, CO_3^{2-}

(3) Sulphate, SO_4^{2-} (4) Sulphide, S^{2-}

Sol: $CH_3COO^- + H_2SO_4 \longrightarrow \underset{(\text{smell of vinegar})}{CH_3COOH} + HSO_4^-$

Ans: (1)

61. Select the reagents that reduce nitriles to primary amines:

A. (i) $LiAlH_4$; (ii) H_2O

B. $Sn + HCl$

C. H_2 / Ni

D. $Na(Hg) / C_2H_5OH$

E. Br_2 / aq.NaOH

Choose the correct answer from the options given below:

(1) A, B and C only (2) A, C and D only (3) A, D and E only (4) B, D and E only

Sol: $R - CN \rightarrow R - CH_2 - NH_2$

$LiAlH_4 / H_2O, H_2 / Ni, Na(Hg) / C_2H_5OH$

Ans: (2)

62. Identify the ***incorrect*** statement from the following:

(1) Carbon has the ability to form $p\pi - p\pi$ multiple bond with itself.

(2) ECl_3 (E = B and Al) is a monomer when E = B and a dimer when E = Al.

(3) Oxygen exhibits only -2 oxidation state.

(4) The order of catenation property of Group 14 elements is $C >> Si > Ge \approx Sn$.

Sol: oxygen exhibits oxidation state $-2, 0, \text{to} + 2, -1, -\frac{1}{2}$

Ans: (3)

63. Although $+3$ oxidation state is most common in lanthanoids, cerium still shows $+4$ oxidation state because:

(1) Its nearest inert gas is Radon.

(2) After losing one more electron, it acquires $4f^{14}$ electronic configuration.

(3) Its atomic number is 61.

(4) After losing one more electron, it acquires $4f^0$ electronic configuration.

Sol: $Ce^{4+} : [Xe]4f^0$

Ans: (4)

64. During Lasagne's test, the elements present in an organic compound are converted from:

(1) covalent form to covalent form (2) ionic form to ionic form

(3) covalent form to ionic form (4) ionic form to covalent form

Sol: Covalent form to ionic form

Ans: (3)

65. The number of hydrogen atoms present in 5.4 g of urea is:

(Given: Molar mass of urea: 60 g mol^{-1}, N_A : 6.022×10^{23} particles mol^{-1})

(1) 2.168×10^{23} (2) 2.168×10^{22} (3) 1.084×10^{22} (4) 1.084×10^{23}

Sol: NH_2CONH_2 : 1 mol contains 4 mol 'H' atoms

$60g \rightarrow 4\times6.022\times10^{23}$ atoms

$5.4g \rightarrow ?$

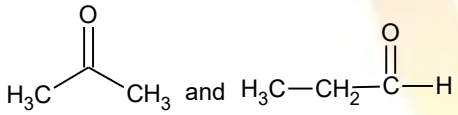
Number of atoms = $\frac{4\times6.022\times10^{23}\times5.4}{60} = 2.168\times10^{23}$ atoms

Ans: (1)

66. The pair of molecules that are metamers among the following is:

(1) $CH_3CH_2CH_2OH$ and $CH_3-CH(OH)-CH_3$

(2) $CH_3OCH_2CH_2CH_3$ and $CH_3CH_2OCH_2CH_3$

(3)  and

(4) $CH_3CH_2CH_2CH_2CH_3$ and $(CH_3)_2CHCH_2CH_3$

Sol: Metamers differ in the carbon chain on either side of functional group Metamers differ in the carbon chain on either side of functional group

Ans: (2)

67. Identify the **incorrect** statement from the following:

(1) $P(C_2H_5)_3$ and $As(C_6H_5)_3$ form $d\pi-d\pi$ bond with transition metals.

(2) Nitrogen can form $d\pi-p\pi$ bond with oxygen.

(3) Nitrogen can form $p\pi-p\pi$ multiple bonds with itself.

(4) Phosphorus, arsenic and antimony show catenation property.

Sol: Nitrogen cannot form $d\pi-p\pi$ bond with oxygen Because for both there is no d-orbital.

Ans: (2)

68. Phenolphthalein is used as an indicator for the titration of sodium hydroxide solution against a standard solution of oxalic acid. The colour change that is observed at an alkaline pH close to the equivalence point during this titration is:

(1) pinkish red to yellow (2) yellow to pinkish red

(3) colourless to pink (4) pink to colourless

Sol: Colour of phenolphthalein in alkaline medium = pink and neutral & acid medium is colourless

OR

$$\underset{\text{weak acid}}{H_2C_2O_4} + \underset{\text{st.base}}{NaOH} \longrightarrow \underset{\text{basic salt}}{Na_2C_2O_4} + H_2O$$

$\therefore$ pH at equivalence point is in alkaline range $H_2C_2O_4$ is added from Burnette, the end point is pink to colourless.

Ans: (4)

69. Match **List I** with **List II**:

	List I		**List II**
A.	C_2H_4	I.	3 σ bonds, 2 π bonds
B.	C_2H_2	II.	3 σ bonds, one lone pair
C.	CH_4	III.	4 σ bonds
D.	NH_3	IV.	5 σ bonds, 1 π bond

Choose the **correct** answer from the options given below:

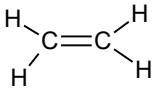
(1) A-IV, B-I, C-III, D-II

(2) A-III, B-IV, C-II, D-I

(3) A-I, B-II, C-IV, D-III

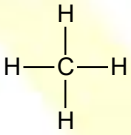
(4) A-II, B-III, C-I, D-IV

Sol:

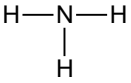


5σ & 1π

3σ & 2π



4σ



3σ & one lone pair

Ans: (1)

70. At a certain temperature, T (K), during a process, 500 J is absorbed by the system and work of 200 J is done by the system. Then change in internal energy of the system is:

(1) 700 J (2) 300 J (3) 400 J (4) 500 J

Sol: $\Delta U = q + \omega$

$= 500 - 200 = 300\,J$

Ans: (2)

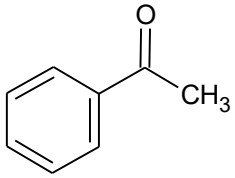
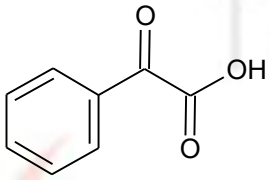
71. Methane reacts with steam at 1273 K in the presence of nickel catalyst to form.

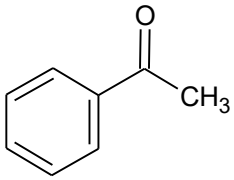
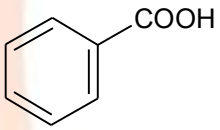
(1) CO and H_2 (2) CO and H_2O (3) CO_2 and H_2O (4) CO_2 and H_2

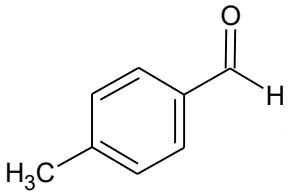
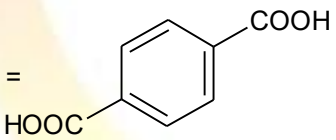
Sol: $CH_4 + H_2O \xrightarrow{Ni} CO + 3H_2$

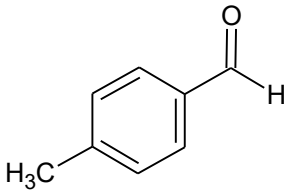
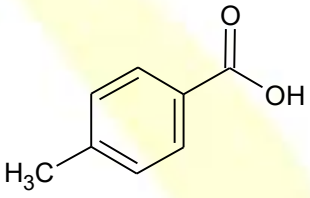
Ans: (1)

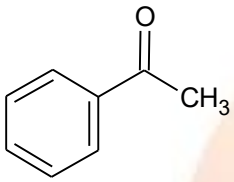
72. Compound $P(C_8H_8O)$ given a red orange precipitate with $2,4-\text{DNP}$ reagent and it does not reduce Fehling's reagent. On drastic oxidation with chromic acid, P gives an aromatic product Q that produces effervescence on treating with $\text{aq.}NaHCO_3$. Compounds P and Q, respectively, are:

(1) P =  ; Q = 

(2) P =  ; Q = 

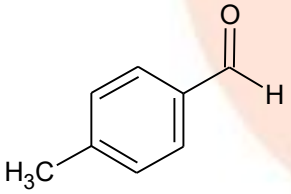
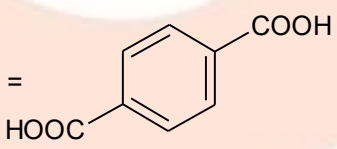
(3) P =  ; Q = 

(4) P =  ; Q = 

Sol: 

Gives red - orange ppt with 2,4 – DNP and does not reduce Fehling's solution on oxidation it gives benzoic acid.

p–methyl benzaldehyde gives negative test for Fehling's solution.

P =  ; Q = 

Ans: (2, 3)

73. A solution of copper sulphate is electrolysed for 10 minutes with a current of 1.5 amperes. The mass of copper deposited at cathode is:

(Given : Molar mass of $Cu = 63\,g\,mol^{-1}$; $1F = 96487\,C\,mol^{-1}$)

(1) 2.4036 g (2) 1.7018 g (3) 0.5876 g (4) 0.2938 g

Sol: $Q = I\,t = 1.5 \times 10 \times 60$ C

$\frac{63}{2}\,g \Rightarrow 96487\,C$

Mass of copper deposited $= \frac{1.5 \times 10 \times 60 \times 31.5}{96487} = 0.2938\,g$

Ans: (4)

74. The functional group that can be identified through phthalein dye test is:

(1) Phenolic (2) Alcohol (3) Aldehyde (4) Carboxylic acid

Sol: It is a characteristic of phenolic group.

Ans: (1)

75. The correct statement with regard to the secondary structure of DNA/RNA is:

(1) DNA possesses a single stand helix structure and contains uracil as one of the four bases.

(2) RNA possesses a single strand helix structure and contains thymine as one of the four bases.

(3) DNA possesses a double strand helix structure and contains thymine as one of he four bases.

(4) RNA possesses a double strand helix structure and contains uracil as one of the four bases.

Sol: DNA possess double helix structure with thymine as one of the four bases. The secondary structure of RNA is single stranded helixs.

Ans: (3)

76. Identify the correct statements:

A. The molality of $2.5\,g$ of ethanoic acid (Molar mass: $60\,g\,mol^{-1}$) in 75 g of benzene solution is $0.556\,m$

B. The molarity of a solution containing 5 g of $NaOH$ (molar mass: $40\,g\,mol^{-1}$) in 450 mL of solution is 0.278 M at 298 K

C. Aquatic species are more comfortable in cold water

D. The solubility of gas increases with decrease in pressure

E. For a binary mixture of A and B are n_A and n_B respectively. The mole fraction of B will be $x_B = \frac{n_A}{n_A + n_B}$

Choose the correct answer from the options given below:

(1) A and C only (2) A, B and C only

(3) A, D and E only (4) A and B only

Sol: A. $\frac{2.5}{60} \times \frac{1000}{75} = 0.556\,m$

B. $\frac{5}{40}\times\frac{1000}{450}=0.278\,\text{m}$

C : Aquatic species are more comfortable in cold water as amount of dissolved oxygen is more at lower temperature.

D : Solubility of gas decreases with decrease in pressure

E : Mole fraction of $B=\frac{n_B}{n_A+n_B}$

Ans: (2)

77. Mixture of chloroform and acetone forms a solution with negative deviation from Raoult's law due to:

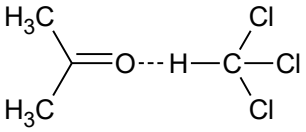
(1) Formation of hydrogen bonding between acetone and chloroform.

(2) Increase in escaping tendency of molecules of each component.

(3) Stronger intermolecular forces between chloroform molecules than those between chloroform and acetone molecules.

(4) Repulsive forces.

Sol:



Formation of $H-$bonding.

Ans: (1)

78. At 298 K, a certain buffer solution contains equal concentrations of X^- and HX, K_b for X^- is 10^{-10} .What is the pH of this buffer solution?

(1) 2 (2) 10 (3) 4 (4) 6

Sol: K_b of $X^-=10^{-10}, K_a$ of $HX=10^{-4}$

$$pH=p^{ka}+\log\frac{\left[X^-\right]}{\left[HX\right]}$$

$$\left[X\right]=\left[HX\right]$$

$$\therefore pH=pK_a=4$$

Ans: (3)

79. Identify the *incorrect* statement from the following:

(1) The IUPAC name of the element with atomic number 107 is Unnilseptium.

(2) The largest and the smallest species among Mg, Mg^{2+}, Al and Al^{3+} are Al and Mg^{2+}, respectively.

(3) The similarity in behaviour of Li with Mg is referred to as 'diagonal relationship'.

(4) The oxidation state and covalency of Al in $\left[AlCl\left(H_2O\right)_5\right]^{2+}$ are 3 and 6 respectively.

Sol: Largest is Mg smallest is Al^{3+}.

Ans: (2)

80. The correct order of increasing metallic character of Na, Be, P, Mg and Si is:

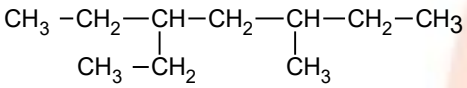
(1) $P < Si < Be < Mg < Na$ (2) $Be < Si < P < Mg < Na$

(3) $P < Si < Na < Mg < Be$ (4) $P < Mg < Be < Si < Na$

Sol: Metallic character increases down the group and decreases across the period.

Ans: (1)

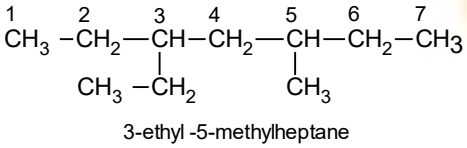
81. The correct IUPAC name of the following compound is:



(1) 2, 4-diethylhexane (2) 3, 5-diethylhexane

(3) 3-ethyl-5-methylheptane (4) 3-methyl-5-ethylheptane

Sol:



3-ethyl -5-methylheptane

Ans: (3)

82. Match List I with List II

List I (Complex/ion)		List II (Shape/geometry)	
A.	$\left[Pt(Cl_2)(NH_3)_2\right]$	I.	Octahedral
B	$\left[Co(NH_3)_6\right]Cl_3$	II.	Trigonal bipyramidal
C.	$[NiCl_4]^{2-}$	III.	Square planar
D.	$\left[Fe(CO)_5\right]$	IV.	Tetrahedral

Choose the correct answer from the options given below:

(1) A-I, B-III, C-IV, D-II (2) A-III, B-IV, C-I, D-II

(3) A-IV, B-I, C-III, D-II (4) A-III, B-I, C-IV, D-II

Sol: A-III, B-I, C-IV, D-II

$\left[Pt(Cl_2)(NH_3)_2\right]$: Square planar

$\left[Co(NH_3)_6\right]Cl_3$: Octahedral

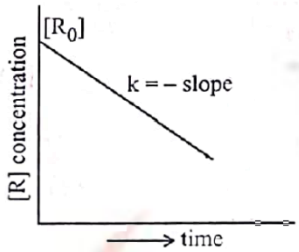
$[NiCl_4]^{2-}$: Tetrahedral

$\left[Fe(CO)_5\right]$: Trigonal bipyramidal

Ans: (4)

83. For a certain reaction $R \rightarrow$ Product, the plot of concentration $[R]$ vs time has a negative slope as shown.

The order of reaction is:



(1) 0 (2) 1 (3) 2 (4) 2.5

Sol: $[R]=[R_0]-kt$ for zero order

$y = C + mx$

$\therefore$ The graph represent zero order reaction.

Ans: (1)

84. Which one of the following is an ambidentate ligand?

(1) Ethylenediaminetetraacetate ion (2) Oxalate

(3) Ethane-1, 2-diamine (4) Thiocyanate

Sol: SCN^- Thiocyanate has S and N as donar atoms. It is an ambidentate ligand.

Ans: (4)

85. Consider the following reaction:

$2A(g)+B(g) \rightarrow 2D(g)$

$\Delta U^{\ominus} = -10 kJmol^{-1}$ and $\Delta S^{\ominus} = -44 JK^{-1}$ at $298K$.

Identify the correct option with $\Delta G^{\ominus}$ for the reaction and spontaneity of the reaction at $298K$.

(Given: $R = 8.31 Jmol^{-1}K^{-1}$)

(1) $-1.635 kJmol^{-1}$, spontaneous (2) $+0.63568 kJmol^{-1}$, non-spontaneous

(3) $-0.63568 kJmol^{-1}$, spontaneous (4) $+1.635 kJmol^{-1}$, non-spontaneous

Sol: $2A(g)+B(g) \rightarrow 2D(g)$

$\Delta n = 2-3 = -1$

$\Delta H = \Delta U + RT \cdot \Delta n$

$\Delta H = -(10 \times 1000) + (8.314 \times 298)(-1)$

$\Delta H = -12,477.5 J$

$\Delta G = \Delta H - T \cdot \Delta S$

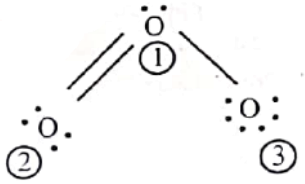
$= -12,477.5 - 298(-44)$

$= -12,477.5 + 13,112 = +634.5 J$

$\Delta G = +ve$ & non-spontaneous

Ans: (2)

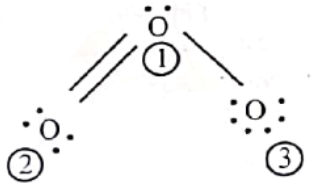
86. Question:



The correct formal charges on oxygen atoms numbered 2, 1 and 3 respectively are:

(1) $-1,0,+1$ (2) $0,+1,-1$ (3) $0,0,0$ (4) $+1,0,-1$

Sol:



Formal Charge on $1=+1,\ 2=0,\ 3=-1$

Ans: (4)

87. Given below are certain reactions. Identify the reaction for which $K_p \neq K_c$.

(1) $H_2(g)+I_2(g) \rightleftharpoons 2HI(g)$ (2) $N_2(g)+O_2(g) \rightleftharpoons 2NO(g)$

(3) $N_2(g)+3H_2(g) \rightleftharpoons 2NH_3(g)$ (4) $H_2O(g)+CO(g) \rightleftharpoons H_2(g)+CO_2(g)$

Sol: If $\Delta n = 0, K_p = K_c$

If $\Delta n \neq 0, K_p \neq K_c$

$N_2(g)+3H_2(g) \rightleftharpoons 2NH_3(g)$

$\Delta n = 2-4 = -2 \neq 0$

$\therefore K_p \neq K_c$

Ans: (3)

88. Given below is an expression for the rate constant of a first order reaction occurring at a certain temperature, $T(K)$.

$$\ln k = 14.34 - \frac{1.25\times10^4}{T}$$

The energy of activation in $kcal\ mol^{-1}$ for the reaction is:

(Given: k in $s^{-1}, R = 1.987 cal\ mol^{-1}K^{-1}$)

(1) 12.42 (2) 14.34 (3) 18.63 (4) 24.84

Sol: $\log_e k = \log_e A - \frac{Ea}{RT}$

$$\log_e k = 14.34 - \frac{1.25\times10^4}{T}$$

$$\Rightarrow \frac{E_a}{R} = 1.25\times10^4$$

$$E_a = 1.25\times10^4 \times R$$

$$E_a = 1.25\times10^4 \times 1.987 = 24.84 kcal$$

Ans: (4)

89. The following two reactions give the same foul smelling product Z.

$$C_2H_5Cl \xrightarrow{X} Z$$

$$C_2H_5\overset{O}{\overset{\|}{C}}NH_2 \xrightarrow{Br_2,\ NaOH} Y \xrightarrow[\Delta]{CHCl_3/\text{ethanolic KOH}} Z$$

X and Z, respectively, are:

(1) $X = AgCN;\ Z = C_2H_5CN$

(2) $X = KCN;\ Z = C_2H_5CN$

(3) $X = KCN;\ Z = C_2H_5NC$

(4) $X = AgCN;\ Z = C_2H_5NC$

Sol:

$$C_2H_5\overset{O}{\overset{\|}{C}}NH_2 \xrightarrow{Br_2,\ NaOH} \underset{(Y)}{C_2H_5-NH_2} \xrightarrow[\Delta]{CHCl_3/\text{ethanolic KOH}} \underset{(Z)}{C_2H_5-\overset{+}{N}\equiv\overset{-}{C}}$$

$$C_2H_5Cl \xrightarrow{AgCN(X)} \underset{(Z)}{C_2H_5\overset{+}{N}\equiv\overset{-}{C}}$$

Ans: (4)

90. Match List I and List II:

List I (Complex)		List II (Type of isomerism)	
A.	$\left[Pt(NH_3)_2Cl_2\right]$	I.	Optical
B	$\left[Co(en)_3\right]^{3+}$	II.	Solvate
C.	$\left[Co(NH_3)_5NO_2\right]Cl_2$	III.	Geometrical
D.	$\left[Cr(H_2O)_6\right]Cl_3$	IV.	Linkage

Choose the correct answer from the options given below:

(1) A-III, B-I, C-II, D-IV

(2) A-I, B-III, C-II, D-IV

(3) A-III, B-I, C-IV, D-II

(4) A-II, B-IV, C-III, D-I

Sol: A-III, B-I, C-IV, D-II

$\left[Pt(NH_3)_2\,Cl_2\right]$- Ma_2b_2, sq. planar show Geometrical Isomerism

$\left[Co(en)_3\right]^{3+}$ - $\left[M(AA)_3\right]$ type show optical isomerism

$\left[Co(NH_3)_5\,NO_2\right]Cl_2$ - has Ambidentate ligand show Linkage isomerism

$\left[Cr(H_2O)_6\right]Cl_3$ - can show hydrate or solvate isomerism

Ans: (3)

Biology

Multiple Choice Questions with one correct answer. A correct answer carries 4 marks. A wrong answer carries a penalty of 1 mark. 45 x 4 = 180

91. "The Evil Quartet" of biodiversity loss includes which of the following?

(1) Over-exploitation; Alien species invasions; Air pollution; Co-extinctions

(2) Habitat loss and fragmentation; over-exploitation; Alien species invasions; Co-extinctions

(3) Habitat loss and fragmentation; Air pollution; Water pollution; Co-extinctions

(4) Over-exploitation; Alien species invasions; Soil pollution; Co-extinctions

Ans: (2)

92. Which one of the following is the site for active ribosomal RNA synthesis?

(1) Nucleolus (2) Chromatin (3) Centrosome (4) Kinetochore

Ans: (1)

93. Match List I with List II

	List – I (Phase of cell)		**List II (Activity)**
A	G_1 phase	I	Actula cell division occurs
B	S phase	II	Cell is metabolically active and continuously gros but does not replicate its DNA
C	G2 phase	III	Synthesis of DNA occurs and the amount of DNA per cell doubles
D	M phase	IV	Proteins are synthesized while cell growth continues

choose the correction answer from the options given below:

(1) A – II, B – III, C – IV, D - I (2) A – jIII, B - IV, C – I, D – II

(3) A – I, B – II , C – III, D – IV (4) A – IV, B – I, C – II, D - III

Ans: (1)

94. Match List I with List II:

	List – I (Phase of cell)		**List II (Activity)**
A	Productivity	I	Gross primary productivity minus respiration losses
B	Net primary productivity	II	Rate of formation of new organic matter by consumers
C	Gross primary productivity	III	Rate of biomass production
D	Secondary productivity	IV	Rate of production of organic matter during photosynthesis

Choose the **correct** answer from the options given below:

(1) A-I, B-II, C-III, D-IV (2) A-III, B-I, C-IV, D-II

(3) A-III, B-I, C-II, D-IV (4) A-I, B-III, C-IV, D-II

Ans: (2)

95. Which of the following statements are correct?

A. The Amazon rainforest being cut and cleared for cultivation of soyabeans is an example of habitat loss.

B. Steller's sea cow and passenger pigeon became extinct due to over-exploitation by humans.

C. The Nile perch introduced into Lake Victoria in East Africa helped in population growth of cichlid fish in the lake.

D. Water hyacinth is an invasive species.

E. When a species becomes extinct, the plant and animal species associated with it are not affected.

Choose the **correct** answer from the options given below:

(1) B, C and D only (2) A, B and D only

(3) A, B and E only (4) C, D and E only

Ans: (2)

96. Identify the correct statements about biomolecules.

A. Lipids are generally water soluble. B. Proteins are polypeptides.

C. Polysaccharides are long chains of sugars.

D. Adenine and guanine are substituted pyrimidines.

E. Almost all enzymes are proteins.

Choose the **correct** answer from the options given below:

(1) C, D and E only (2) B, C and E only

(3) B, D and E only (4) A, B and C only

Ans: (2)

97. How many ATP and NADPH molecules are required to make one molecule of glucose through the Calvin pathway?

(1) 18 ATP and 12 NADPH (2) 6 ATP and 12 NADPH

(3) 24 ATP and 18 NADPH (4) 12 ATP and 18 NADPH

Ans: (1)

98. Which of the following statements are **not** true regarding restriction endonucleases?

A. They are called molecular scissors.

B. These are the enzymes responsible for restricting the growth of bacteriophages in *E. coli*.

C. They cut the DNA only at the centre of the palindromic sites.

D. They remove nucleotides only from the ends of DNA fragments.

E. They recognise specific palindromic base-pair sequences.

Choose the answer from the options given below:

(1) A and B only (2) D and E only

(3) C and D only (4) A and E only

Ans: (3)

99. Match List I with List II:

	List – I		**List II**
A	Decomposition	I	Accumulation of dark coloured amorphous colloidal substance
B	Detritus	II	Release of inorganic nutrients by the activity of microbes in soil
C	Mineralisation	III	Breaking down of complex organic matter into inorganic substances
D	Humification	IV	Dead remains of plants and animals including fecal matter

Choose the **correct** answer from the options given below:

(1) A-I, B-II, C-III, D-IV (2) A-IV, B-III, C-I, D-II

(3) A-III, B-IV, C-II, D-I (4) A-III, B-II, C-I, D-IV

Ans: (3)

100. In which one of the following, the ovules are **not** enclosed by an ovary wall and remain exposed?

(1) *Selaginella* (2) *Funaria* (3) *Pinus* (4) *Wolffia*

Ans: (3)

101. Match List I with List II:

	List – I (Placentation)		**List – II (Example)**
A	Marginal	I	Mustard
B	Axile	II	Pea
C	Parietal	III	Marigold
D	Basal	IV	Lemon

Choose the **correct** answer from the options given below:

(1) A-I, B-III, C-II, D-IV (2) A-IV, B-II, C-I, D-III

(3) A-II, B-IV, C-I, D-III (4) A-III, B-I, C-IV, D-II

Ans: (3)

102. In angiosperms, root hairs arise from which one of the following regions of the root?

(1) The root cap zone
(2) The region of meristematic activity
(3) The region of elongation
(4) The region of maturation

Ans: (4)

103. Which one of the following is not a characteristic of plant cells in the phase of elongation?

(1) Increased vacuolation
(2) Large conspicuous nuclei
(3) Cell enlargement
(4) New cell wall deposition

Ans: (2)

104. Which of the following statements are correct with reference to a transcription unit?

A. A transcription unit in DNA is defined primarily by three regions: promoter, structural gene and terminator.

B. The promoter is said to be located towards the 5'-end of the structural gene.

C. The promoter is a DNA sequence that provides binding site for RNA polymerase.

D. The promoter defines the template and coding strands.

E. The terminator is located towards the 3'-end of the coding strand and it defines the end of the process of transcription.

Choose the **correct** answer from the options given below:

(1) A, B, C, D and E
(2) B, C, D and E only
(3) A, C, D and E only
(4) A, B, C and D only

Ans: (1)

105. Alpha-helix is found in which level of protein structure?

(1) Quaternary structure
(2) Tertiary structure
(3) Primary structure
(4) Secondary structure

Ans: (4)

106. Which of the following statements are **correct** regarding amino acids?

A. They are substituted methanes.

B. Serine is an aromatic amino acid.

C. Valine is a neutral amino acid.

D. Lysine is an acidic amino acid.

Choose the **correct** answer from the options given below:

(1) C and D only
(2) A and B only
(3) A and C only
(4) B and C only

Ans: (3)

107.The main function of bulliform cells in grasses is:

(1) to make the leaf impermeable to fungal spores.

(2) to perform photosynthesis.

(3) to minimize water loss during water stress.

(4) to transport water.

Ans: (3)

108.Find the **incorrect** statement(s) about photosynthesis from the following:

A. The water splitting complex is associated with PS I.

B. C_4 plants use the C_3 pathway of CO_2 fixation as the main biosynthetic pathway.

C. In C_4 plants, photorespiration does not occur.

D. C_3 plants exhibit 'Kranz' anatomy.

E. ATP synthesis in chloroplast occurs through chemiosmosis.

Choose the answer from the options given below:

(1) B only (2) A and D only (3) B and C only (4) B and E only

Ans: (2)

109.Match List I with List II:

	List I		**List II**
A	Conjunctive tissue	I	Specialised cells in the vicinity of guard cells
B	Casparian strips	II	Endodermal cells rich in starch
C	Subsidiary cells	III	Tissue between xylem and phloem
D	Starch sheath	IV	Endodermal cells with suberin deposition

Choose the **correct** answer from the options given below:

(1) A-IV, B-III, C-I, D-II (2) A-III, B-IV, C-II, D-I

(3) A-III, B-IV, C-I, D-II (4) A-IV, B-III, C-II, D-I

Ans: (3)

110.Match List I with List II:

	List I		**List II**
A	Genetically modified organism	I	Agrobacterium tumefaciens
B	Thermostable DNA polymerase	II	Bt cotton
C	Ti plasmid	III	Thermus aquaticus
D	pBR322	IV	Escherichia coli

Choose the **correct** answer from the options given below:

(1) A-II, B-I, C-IV, D-III (2) A-I, B-IV, C-III, D-II

(3) A-II, B-III, C-I, D-IV (4) A-I, B-II, C-IV, D-III

Ans: (3)

111. Heterophyllous development in response to environment is an example of which of the following phenomena?

(1) Dedifferentiation
(2) Elasticity
(3) Redifferentiation
(4) Plasticity

Ans: (4)

112. In racemose inflorescence, _____.

(1) the main axis terminates in a flower
(2) the growth is limited
(3) flowers are borne in an acropetal succession
(4) flowers are solitary

Ans: (3)

113. Which one of the following disorders is caused by the substitution of Glutamic acid (Glu) by Valine (Val) at the sixth position of the beta globin chain of the haemoglobin molecule?

(1) Haemophilia
(2) Thalassemia
(3) Sickle-cell anaemia
(4) Phenylketonuria

Ans: (3)

114. Match List I with List II:

	List I		**List II**
A	Incomplete dominance	I	Human skin colour
B	Co-dominance	II	Inheritance of flower colour in Antirrhinum sp.
C	Pleiotropy	III	Phenylketonuria disease in humans
D	Polygenic inheritance	IV	ABO blood groups

Choose the **correct** answer from the options given below:

(1) A-II, B-IV, C-III, D-I
(2) A-I, B-III, C-II, D-IV
(3) A-II, B-I, C-III, D-IV
(4) A-I, B-IV, C-III, D-II

Ans: (1)

115. Arrange the following in the correct developmental sequence related to microsporogenesis:

A. Microspore tetrads
B. Sporogenous tissue
C. Pollen grains
D. Pollen mother cells

Choose the **correct** answer from the options given below:

(1) A, C, B, D
(2) B, D, C, A
(3) B, D, A, C
(4) A, D, C, B

Ans: (3)

116. Arrange the following steps of DNA fingerprinting in a correct sequence.

A. Isolation of DNA and its digestion by restriction endonucleases.

B. Hybridisation using a labelled VNTR probe.

C. Transferring of separated DNA fragments to synthetic membranes.

D. Detection of hybridised DNA fragments by autoradiography.

E. Separation of DNA fragments by electrophoresis.

Choose the **correct** answer from the options given below:

(1) A, E, C, B, D

(2) A, E, B, C, D

(3) A, B, D, C, E

(4) A, D, B, E, C

Ans: (1)

117. Exploring molecular, genetic and species-level diversity for products of economic importance is called:

(1) Biomagnification

(2) Biofortification

(3) Bioremediation

(4) Bioprospecting

Ans: (4)

118. Which of the following statements are true with reference to the sex-determination in honeybees?

A. An offspring formed from the union of a sperm and an egg, develops as a female (queen or worker).

B. An unfertilized egg develops as a male by parthenogenesis.

C. A male has half the number of chromosomes than that of a female.

D. Males produce sperms by meiosis.

E. Honeybees have a haplodiploid sex-determination system.

Choose the **correct** answer from the options given below:

(1) B, C, D and E only

(2) A, B, C and D only

(3) A, B, D and E only

(4) A, B, C and E only

Ans: (4)

119. Identify the **correct** sequence of steps in each cycle of Polymerase Chain Reaction:

(1) Denaturation $\rightarrow$ Annealing $\rightarrow$ Extension

(2) Denaturation $\rightarrow$ Extension $\rightarrow$ Annealing

(3) Extension $\rightarrow$ Annealing $\rightarrow$ Denaturation

(4) Annealing $\rightarrow$ Denaturation $\rightarrow$ Extension

Ans: (1)

120. Which of the following statements are correct with respect to DNA separation, isolation and visualization?

A. The cutting of DNA is done by molecular scissors.

B. The DNA fragments separate according to their size in an agarose gel, upon electrophoresis.

C. The separated DNA fragments can be seen without staining when exposed to UV light.

D. The separated DNA fragments, when stained with ethidium bromide, can be seen in visible light.

Choose the **correct** answer from the options given below:

(1) A and D only (2) B and D only

(3) B and C only (4) A and B only

Ans: (4)

121. The main criteria used for five Kingdom Classication proposed by R.H Whittaker (1969) included:

A. A. Cell structure

B. Body organization.

C. Presence of flagellum

D. Reproduction

E. Phylogenetic relationships

Choose the correct answer from the options given below :

(1) A, B, D and E only (2) A, B, C, D and E

(3) A, B and E only (4) B, C and D only

Ans: (1)

122. Which one of the following is a triploid cell

(1) Central cell (2) Primary endosperm cell

(3) Zygote (4) Synergid

Ans: (2)

123. Which of the following statements are correct with reference to packaging of DNA helix ?

A. Histones are organized to form a unit of eight molecules called histone octamer.

B. Histones are negatively charged basic proteins.

C. Histones are rich in the basic amino acid residues - lysine and argininc.

D. The positively charged DNA is wrapped around the histone octamer to form nucleosome.

E. The packaging of chromatin at higher levels requires an additional set of proteins called non-histone chromosomal proteins.

Choose the correct answer from the options given below :

(1) A, B and D only (2) A, C and E only (3) C, D and E only (4) B, D and E only

Ans: (2)

124. Which of the following is an in situ conservation method ?

(1) Sacred Groves
(2) Wildlife Safari Parks
(3) Botanical Gardens
(4) Seed Banks

Ans: (1)

125. In the lac operon, the z gene codes for :

(1) transacety lase
(2) the repressor of lae operon
(3) permease
(4) beta-galactosidase

Ans: (4)

126. Match List I with List II :

List I (Growth Regulator)		List II (Function/Effect)	
A	2,4-D	I.	Brewing industry
B.	GA_3	II.	Stimulation of stomatal closure
C.	Kinetin	III.	Herbicide
D.	ABA	IV.	Nutrient mobilisation

Choose the correct answer from the options given below :

(1) A-IV, B-III, C-II, D-I
(2) A-I, B-II, C-IV, D-III
(3) A-III, B-I, C-IV,D-II
(4) A-I,B-IV,C-III, D-II

Ans: (3)

127. Arrange the following steps somatic hybridisation in a correct sequence.

A. Digestion of cell walls.

B. Isolation of naked protoplasts.

C. Fusion of protoplasts to get hybrid protoplast.

D. Isolation of single cells from two different varieties of plants.

E. Growing of hybrid protoplast to form a new plant.

Choose the correct answer from the options given below :

(1) E,A, B, C, D
(2) D, A, B, C,E
(3) E.B, A, D, C
(4) D, B, A, E, C

Ans: (2)

128. $2\left(C_{51}H_{98}O_6\right) + 145\, O_2 \rightarrow 102\, CO_2 + 98\, H_2O + energy$

The Respiratory Quotient (RQ) of a biomolecule used for respiration, as per the above equation, would be :

(1) Less than 0.5
(2) Between 0.5 and 0.95
(3) Between 1.25 and 2
(4) 1.0

Ans: (2)

129. Since the origin and diversification of life on Earth, there have been five episodes of mass extinction of species. How is the sixth extinction, which is in progress, different from the previous episodes ?

(1) The current species extinction rates are far lower than those in previous episodes.

(2) The present species extinction rates are 100 to 1000 times daster than in the pre-human times.

(3) The present net species extinetion rate is zero.

(4) The current species extinction rate is nearly 10 times faster than that in previous episodes.

Ans: (2)

130.Match List I with List II :

List I		List II	
A.	Trypsin	I.	Intercellular ground substance
B.	Morphine.	II.	Lectin
C.	Concanavalin	III.	Enzyme
D.	Collagen	IV.	Alkaloid

Choose the **correct** answer from the options given below:

(1) A-III, B-IV, C-II, D-I

(2) A-I, B-II, C-III, D-IV

(3) A-III, B-II, C-IV, D-I

(4) A-IV, B-III, C-II, D-I

Ans: (1)

131.Which one of the following statements is not true about the universal rules of binomial nomenclature ?

(1) Both the words in a biological name, when handwritten, are separately underlined or printed in italics.

(2) The specific epithet in the biological name starts with a small letter.

(3) The first word in the biological name represents the specific epithet, while the second component denotes the genus.

(4) Biological names are generally in Latin.

Ans: (3)

132.The enzyme required for carboxylation in the Calvin cycle is :

(1) PEP carboxylase

(2) RuBP carboxylase - oxygenase

(3) Carboxypeptidase

(4) Hexokinase

Ans: (2)

133. Which of the following floral formula is the correct floral formula of Solanaceae family ?.

(1) $\oplus ⚥ K_{(5)} \overset{\frown}{C_{(5)} A_5} \underline{G}_{(2)}$

(2) $\oplus ⚥ K_5 \overset{\frown}{C_{(5)} A_5} \underline{G}_{(2)}$

(3) $\oplus ⚥ K_{(5)} C_{(5)} A_5 \underline{G}_{(2)}$

(4) $\oplus ⚥ K_5 C_5 A_5 \underline{G}_{(2)}$

Ans: (1)

134. Which one of the following types of pollination brings genetically different types of pollen grains to the stigma ?

(1) Geitonogamy (2) Autogamy

(3) Xenogamyo (4) Cleistogamy

Ans: (3)

135. Match List I with List II :

	List I (Process)		List II (Location)
A	Glycolysis	I	Inner mitochondrial membrane
B	ETS	II	Mitochondrial matrix
C	Accumulation of protons	III	Cytoplasm
D	Krebs' cycle	IV	Intermembrane space

Choose the **correct** answer from the options given below :

(1) A-I, B-IV, C-III, D-II (2) A-III, B-I, C-IV, D-II

(3) A-IV, B-II, C-I, D-III (4) A-II, B-III, C-IV, D-I

Ans: (2)

136. Insertion of a foreign DNA at BamHI site in an *E. coli* cloning vector pBR322 results in the loss of antibiotic resistance towards :

(1) Gentamycin (2) Ampicillin and tetracycline

(3) Tetracycline (4) Ampicillin

Ans: (3)

137. The sixth mutant codon of beta globin gene causing polymerization of Haemoglobin and change in RBC shape is _____.

(1) CAG (2) GUG (3) AUG (4) GAG

Ans: (2)

138. Choose the correct statement regarding GIFT to overcome infertility.

(1) Ova collected from a female donor are transferred to the uterus of an infertile female.

(2) It is the transfer of an ovum collected from a donor into the fallopian tube of another female who cannot produce ovum but can provide suitable environment for fertilization and development.

(3) Early embryos with up to 8 blastomeres are transferred to the uterus of an infertile female.

(4) Early embryos with up to 8 blastomeres are transferred into the fallopian tube of an infertile female.

Ans: (2)

139. Which one of the following is an appropriate example of 'sexual deceit' ?

(1) Female wasp and fig (2) Cuckoo and crow

(3) *Ophrys* and bumblebee (4) Sea anemone and clown fish

Ans: (3)

140. Evolution of human appears parallel to the progressive development of brain and language skills. As such, the evolution of individual species in the sequence of their appearance is :

(1) *Homo habilis* → *Homo erectus* → *Ramapithecus* → *Neanderthal* → *Homo sapiens*

(2) *Ramapithecus* → *Homo habilis* → *Homo erectus* → *Neanderthal* → *Homo sapiens*

(3) *Homo sapiens* → *Ramapithecus* → *Homo habilis* → *Neanderthal* → *Homo erectus*

(4) *Neanderthal* → *Ramapithecus* → *Homo habilis* → *Homo erectus* → *Homo sapiens*

Ans: (2)

141. Match List I with List II related to embryonic development at various months of pregnancy :

	List I		**List II**
A	The foetus movement starts and hair appears on the head	I	24 weeks of pregnancy
B	The foetus develops limbs and digits	II	20 weeks of pregnancy
C	The foetus develops external genital organs	III	8 weeks of pregnancy
D	The foetus body is covered with fine hair; eyelids separate and eyelashes are formed	IV	12 weeks of pregnancy

Choose the **correct** answer from the options given below :

(1) A-III, B-II, C-IV, D-I

(2) A-II, B-IV, C-III, D-I

(3) A-IV, B-II, C-III, D-I

(4) A-II, B-III, C-IV, D-I

Ans: (4)

142. A group of researchers procured some fish-like animals and upon investigation the following characters were observed :

A. Endoskeleton was made of cartilage.

B. Ectoparasitic; as they were found attached on fish skin with their circular sucking mouth.

C. Paired fins and scales were absent, but 7 pairs of gill slits were present.

Which of the following species of animals did they consider to fit best with these characters ?

(1) *Exocoetus* sp.

(2) *Branchiostoma* sp.

(3) *Petromyzon* sp.

(4) *Scoliodon* sp.

Ans: (3)

143. Spermatogonia undergo a series of cell divisions to produce sperms. Select the **correct** statements from the following :

A. Spermatogonia always undergo meiotic cell division.

B. Primary spermatocytes divide mitotically to produce secondary spermatocytes.

C. Secondary spermatocytes, through their second meiotic division, produce haploid spermatids.

D. Spermatids produce spermatozoa through mitosis.

E. Spermatids transform into spermatozoa by spermiogenesis.

Choose the **correct** answer from the options given below :

(1) C and E only (2) A, C and E only

(3) B, C and D only (4) A and E only

Ans: (1)

144. What is the probability of having children with 'O' blood group, where both mother and father are heterozygous for 'A' and 'B' blood group, respectively ?

(1) 50% (2) 0% (3) 75% (4) 25%

Ans: (4)

145. Arrange the following events occurring in Renin-Angiotensin mechanism in the correct order :

A. Increase in blood pressure and Glomerular filtration rate.

B. Reabsorption of Na^+ and water from distal parts of tubule due to Aldosterone.

C. Fall in Glomerular filtration rate.

D. Vasoconstriction by Angiotensin II and release of Aldosterone.

E. Renin converts Angiotensinogen into Angiotensin I, followed by Angiotensin II.

Choose the **correct** answer from the options given below :

(1) C, A, B, D, E (2) A, D, B, E, C

(3) A, C, E, B, D (4) C, E, D, B, A

Ans: (4)

146. Match List I with List II :

	List I (Respiratory Volume)		**List II (Capacity in mL)**
A	ERV (Expiratory Reserve Volume)	I	2500 – 3000 mL
B	RV (Residual Volume)	II	500 mL
C	IRV (Inspiratory Reserve Volume)	III	1000 – 1100 mL
D	TV (Tidal Volume)	IV	1100 – 1200 mL

Choose the **correct** answer from the options given below :

(1) A-III, B-I, C-IV, D-II (2) A-I, B-III, C-II, D-IV

(3) A-III, B-IV, C-I, D-II (4) A-I, B-II, C-III, D-IV

Ans: (3)

147. Match List I with List II :

	List I		**List II**
A	Progestasert	I	Barrier made of rubber used by females
B	Multiload 375	II	Oral contraceptive
C	Diaphragm	III	Hormone releasing IUD
D	Saheli	IV	Copper releasing IUD

Choose the **correct** answer from the options given below :

(1) A-III, B-IV, C-I, D-II (2) A-III, B-IV, C-II, D-I

(3) A-IV, B-II, C-I, D-III (4) A-IV, B-III, C-I, D-II

Ans: (1)

148. Non-membrane bound cell organelles found in both prokaryotic and eukaryotic cells are ____

(1) Centrosomes (2) Ribosomes

(3) Lysosomes (4) Mitochondria

Ans: (2)

149. Ecological pyramids represent the relationship between the organisms at different trophic levels and they are generally inverted for :

(1) Pyramid of energy in pond ecosystem (2) Pyramid of biomass in sea

(3) Pyramid of number in grassland (4) Pyramid of biomass in grassland

Ans: (2)

150. The flightless bird with forelimbs modified as paddle-like structures suited for swimming is known as :

(1) *Struthio* (2) *Psittacula* (3) *Neophron* (4) *Aptenodytes*

Ans: (4)

151. Match List I with List II:

	List I (Bioactive molecules)		**List II (Importance)**
A.	Streptokinase	I.	Immunosuppressive agent
B.	Statins	II.	Removal of clots from the blood vessels
C.	Lipases	III.	Blood cholesterol-lowering agent
D.	Cyclosporin A	IV.	Detergent formulations

Choose the correct answer from the options given below:

(1) A-II, B-III, C-IV, D-I (2) A-IV, B-III, C-II, D-I

(3) A-II, B-III, C-I, D-IV (4) A-III, B-II, C-IV, D-I

Ans: (1)

152. Choose the correct statements regarding cell organelles and their inclusions.

A. The endomembrane system includes Golgi complex, endoplasmic mitochondria.

B. Rough endoplasmic reticulum bears ribosomes on its surface.

C. Both mitochondria and plastids have circular DNA.

D. A network of microtubules, microfilaments and intermediate filaments present in the cytoplasm is called cytoskeleton.

E. Mitochondrion is a single membrane-bound structure.

Choose the correct answer from the options given below:

(1) C, D and E only (2) A and B only (3) A, B and C only (4) B, C and D only

Ans: (4)

153.Select the set of fishes which belong to the class Osteichthyes:

(1) Devil fish, Cuttlefish and Hagfish

(2) Starfish, Hagfish and Cuttlefish

(3) Flying fish, Angel fish and Fighting fish

(4) Saw fish, Fighting fish and Dog fish

Ans: (3)

154.In a population of a grasshopper species, the chromosome number of some members is 23 and some other members possess 24 chromosomes.

The 23 and 24 chromosome-bearing members in this species are _____.

(1) all males

(2) all females

(3) females and males, respectively

(4) males and females, respectively

Ans: (4)

155.The WBC count of a person's blood sample is 8000/cu.mm. How many eosinophils and lymphocytes would be in the same blood sample approximately?

(1) 160 - 240/cu.mm and 1600 - 2000/cu.mm, respectively

(2) 100 - 120/cu.mm and 160 - 200/cu.mm, respectively

(3) 300 - 500/cu.mm and 500 - 700/cu.mm, respectively

(4) 300 - 500/cu.mm and 1200- 1500/cu.mm, respectively

Ans: (1)

156.The toxin proteins isolated from *Bacillus thuringiensis*, coded by which of the following genes would control cotton bollworms and corn borer, respectively?

(1) cryIAc and cryIIIAb

(2) cryIAc and cryIIAb

(3) cryIAc and cryIAb

(4) cryIIAb and cryIAc

Ans: (3)

157.Match List I with List II:

	List I (Drug)		**List II (Effect)**
A.	Nicotine	I.	Causes sense of euphoria and increased energy
B.	Morphine	II.	Stimulates adrenal gland to release catecholamines into blood circulation
C.	Heroin	III.	Effective sedative and painkiller
D.	Cocaine	IV.	A depressant; slows down body function

Choose the correct answer from the options given below:

(1) A-III, B-II, C-IV, D-I

(2) A-II, B-III, C-IV, D-I

(3) A-II, B-III, C-I, D-IV

(4) A-III, B-II, C-I, D-IV

Ans: (2)

158. Match List I with List II related to muscular/skeletal system:

	List I		List II
A.	Tetany	I.	Inflammation of joints
B.	Arthritis	II.	Autoimmune disorder affecting neuromuscular junction
C.	Myasthenia gravis	III.	Wild contraction in muscle due to low Ca^{++} in body fluid
D.	Muscular dystrophy	IV.	Progressive degeneration of skeletal muscle

Choose the correct answer from the options given below:

(1) A-III, B-I, C-II, D-IV (2) A-IV, B-III, C-II, D-I

(3) A-I, B-II, C-III, D-IV (4) A-III, B-II, C-I, D-IV

Ans: (1)

159. In which animal do haploid cells divide mitotically to produce gametes?

(1) Male honeybees (2) Male grasshoppers (3) Male earthworms (4) Male frogs

Ans: (1)

160. In humans, respiration occurs in the following steps. Arrange these steps in the correct order.

A. Diffusion of O_2 and CO_2 between blood and tissues

B. Diffusion of O_2 and CO_2 across alveolar membrane

C. Pulmonary ventilation by which atmospheric air is drawn in and CO_2 rich alveolar air is released out

D. Cellular respiration

E. Transport of gases by the blood

Choose the correct answer from the options given below:

(1) A, B, C, D, E (2) E, A, C, D, B (3) C, A, B, E, D (4) C, B, E, A, D

Ans: (4)

161. Arrange the following cell layers/structures around the female gamete, from outer to inner side:

(A) Zona pellucida

(B) Perivitelline space

(C) Corona radiata

(D) Plasma membrane of ovum

(1) C, A, D, B (2) C, A, B, D (3) D, B, A, C (4) A, C, B, D

Choose the correct answer from the options given below:

Ans: (2)

162. The human protein named α-1-antitrypsin, obtained from transgenic animals, is used for the treatment of _____.

(1) Alzheimer's disease (2) Emphysema

(3) Rheumatoid arthritis (4) Cystic fibrosis

Ans: (2)

163.Select the correct statements regarding cell membrane in eukaryotic cell.

A. Membrane of human RBCs has approximately 52% protein.

B. Major phospholipids are arranged in a bilayer.

C. Extensions of the plasma membrane into the cell form mesosomes.

D. Tails towards the inner part of lipids are hydrophobic and thus protected from aqueous medium.

E. Glycocalyx is present on the outer surface of the plasma membrane.

Choose the correct answer from the options given below:

(1) A, C and E only (2) B, C and E only (3) C, D and E only (4) A, B and D only

Ans: (4)

164.Male frogs can be distinguished from female frogs due to the presence of:

A. Bulging eyes

B. Vocal sacs

C. Webbed digits in feet

D Copulatory pad on first digit of fore limbs

E. Olive green-coloured skin with dark irregular spots

Choose the correct answer from the options given below:

(1) B and D only (2) B and C only (3) A and B only (4) Cand E only

Ans: (1)

165.Which of the following equations depicts Verhulst-Pearl logistic population growth?

(1) $\frac{dN}{dt} = rN\left(\frac{K-N}{K}\right)$ (2) $\frac{dN}{dt} = rN\left(\frac{K+N}{K}\right)$ (3) $\frac{dN}{dt} = rN\left(\frac{K}{K-N}\right)$ (4) $\frac{dN}{dt} = rN\left(\frac{K-N}{N}\right)$

Ans: (1)

166.Choose the correct statements regarding frog's anatomy:

A. Hepatic poftal system is the special venous connection between liver and intestine.

B. There are (twelve pairs of cranial nerves arising from the brain.

C. The ureters and oviducts open separately into the cloaca in female frogs.

D. Hind-brain consists of cerebellum, medulla oblongata and optic lobes.

E. Sinus venosus joins the right atrium of heart.

Choose the correct answer from the options given below:

(1) B and D only (2) A, C and E only (3) A, B and C only (4) B and C only

Ans: (2)

167.Select the incorrect statements with reference to Rh grouping.

A. Erythroblastosis foetalis is a condition observed having foetus with Rh^{-ve} blood and mother with Rh^{-ve} blood.

B. Rh antigen is observed on RBCs in the majority of human beings.

C. Before blood transfusion, Rh group should also be matched.

D. Rh incompatibility is observed when a pregnant mother is Rh^{-ve} and the foetus is Rh^{+ve}

E. Erythroblastosis foetalis can be avoided by administering anti-Rh antibodies to the mother immediately after the delivery of the second child.

Choose the answer from the options given below:

(1) A and E only (2) A and B only (3) B and C only (4) C and D only

Ans: (1)

168. Which of the following statements are correct with reference to human endoskeleton?

A. Human skull is monocondylic.

B. The joint between any two adjoining vertebrae is a cartilaginous joint.

C. In human beings, the number of cervical vertebrae is seven.

D. All ribs except the last 2 pairs are bicephalic.

E. The occipital bone of skull is articulated with atlas vertebra.

Choose the correct answer from the options given below:

(1) A, B and D only (2) B and E only (3) B, C and E only (4) C, D and E only

Ans: (3)

169. Match List I with List II

	List I		**List II**
A.	Cortisol	I.	Stimulates the formation of alveoli in mammary glands
B.	Aldosterone	II.	Produces anti-inflammatory reactions
C.	Cholecystokinin	III.	Stimulates reabsorption of Na^+ and water from renal tubule
D.	Progesterone	IV.	Stimulates secretion of pancreatic enzymes bile juice

Choose the correct answer from the options given below:

(1) A-III, B-II, C-IV, D-I (2) A-II, B-III, C-IV, D-I

(3) A-IV, B-II, C-I, D-III (4) A-II, B-III, C-I, D-IV

Ans: (2)

170. The following are the stages of life cycle of *Plasmodium*. Arrange the stages in the proper order.

A. The parasites reproduce asexually in RBCs, bursting the cells.

B. The parasites reproduce asexually in liver cells, bursting the cells and releasing into blood.

C. Gametocytes develop in RBCs.

D. Sporozoites reach the liver through the blood.

E. Female mosquito injects sporozoites into humans during bite.

Choose the correct answer from the options given below:

(1) A, B, C, D, E (2) E, D, B, A, C (3) C, A, B, D, E (4) E, C, D, B, A

Ans: (2)

171.Select the incorrect statements from the following:

A. Digestive system in Platyhelminthes is incomplete.

B. Bilateral symmetry is a characteristic feature of adult Echinoderms.

C. Pseudocoelom is possessed by Aschelminthes.

D. Notochord is persistent throughout life in the class Chondrichthyes.

E. Members of class Reptilia maintain a constant body temperature.

Choose the answer from the options given below:

(1) B and E only (2) C and D only (3) A and C only (4) B and D only

Ans: (1)

172.The specific receptors for neurotransmitters in a synapse are present on _____.

(1) Post-synaptic membrane (2) Pre-synaptic membrane

(3) Myelin sheath (4) Schwann cell

Ans: (1)

173.Choose the correct statements regarding muscle contraction.

A. A motor neuron carries a signal sent by the Central Nervous System (CNS) to the sarcolemma of the muscle fibre.

B. The neural signal generates an action potential which causes the release of Ca^{++} into sarcoplasm.

C. Increase in Ca^{++} inactivates the actin for breaking cross bridges.

D. Actin binds to the myosin head to form a cross bridge.

E. Shortening of sarcomere takes place, by pulling actin filaments towards the centre of 'A' band.

Choose the correct answer from the options given below:

(1) A, B, D and E only (2) C and D only

(3) C and E only (4) A and B only

Ans: (1)

174.Which of the following is *not* an example of convergent evolution?

(1) Eyes of octopuses and mammals

(2) Fore limbs of whales and bats

(3) Wings of butterflies and birds

(4) Flippers of penguins and dolphins

Ans: (2)

175.The JGA (Juxta Glomerular Apparatus) is a special sensitive region formed by cellular modifications in _____ related to the same nephron.

(1) Distal convoluted tubule and efferent renal arteriole

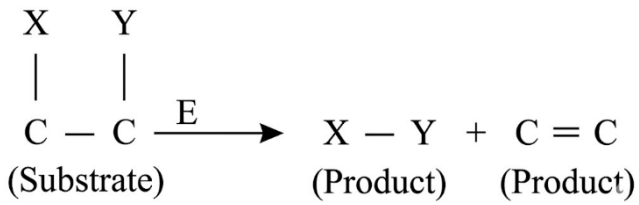
(2) Proximal convoluted tubule and afferent renal arteriole

(3) Distal convoluted tubule and afferent renal arteriole

(4) Proximal convoluted tubule and efferent renal arteriole

Ans: (3)

176.The following reaction depicts the activity of a particular class of enzymes:



Identify the enzyme class 'E' from the following options:

(1) Ligases (2) Lyases (3) Isomerases (4) Transferases

Ans: (2)

177.Match List I with List II:

	List I		List II
A	Molluscs	I	Pulmonary respiration only
B	Reptiles	II	Branchial respiration
C	Adult amphibians	III	Cellular respiration
D	Amoeba	IV	Pulmonary and Cutaneous respiration

Choose the correct answer from the options given below:

(1) A-III, B-II, C-I, D-IV (2) A-II, B-I, C-IV, D-III

(3) A-II, B-I, C-III, D-IV (4) A-I, B-II, C-IV, D-III

Ans: (2)

178.What is the reason behind production of large holes in 'Swiss Cheese'?

(1) The production of large amount of CO_2 by *Clostridium butylicum*

(2) The production of large amount of CO_2 and H_2 by *Trichoderma polysporum*

(3) The production of large amount of CO_2 and H_2 by lactic acid bacteria called *Lactobacillus*

(4) The production of large amount of CO_2 by *Propionibacterium sharmanii*

Ans: (4)

179.Match List I with List II with respect to chronology of evolution of life forms:

	List I		List II
A	About 65 mya	I	Jawless fish probably evolved
B	About 500 mya	II	The dinosaurs suddenly disappeared from the earth
C	About 350 mya	III	Seaweeds and few plants probably existed
D	About 320 mya	IV	Invertebrates were formed and became active

Choose the correct answer from the options given below:

(1) A-II, B-IV, C-I, D-III (2) A-II, B-IV, C-III, D-I

(3) A-I, B-II, C-III, D-IV (4) A-III, B-IV, C-I, D-II

Ans: (1)

180. Choose the correct statements regarding population interactions between two species.

A. In both parasitism and commensalism, only one species benefits and the other species is harmed.

B. Both species benefit in mutualism.

C. Both species benefit in commensalism.

D. In parasitism, only one species benefits and the other species is harmed.

E. In amensalism, one species is harmed and the other is unaffected.

Choose the correct answer from the options given below:

(1) A and D only (2) A and B only (3) B and E only (4) B, D and E only

Ans: (4)