

Topics: Metals and Non-metals

Subtopics: Reaction between Metals and Non-metals

Questions

- Q1. Write the electron dot structures for sodium, oxygen and magnesium.
- Q2. Show the formation of Na_2O and MgO by the transfer of electrons. What are ions present in these compounds?
- Q3. Why do ionic compounds have high melting points ?
- Q4. Write the electron dot structures for potassium and chlorine.
- Q5. Show the formation of KCl by the transfer of electrons.
- Q6. (a) Show the formation of CaS by the transfer of electrons.
(b) Name the ions present in this compound CaS .
- Q7. Explain the formation of ionic compound CaO with electron dot structure.
- Q8. Why are ionic compounds usually hard?
- Q9. How is it that ionic compounds in the solid state do not conduct electricity but they do so when in molten state?
- Q10. Why are aqueous solutions of ionic compounds able to conduct electricity?