

Topics: Metals and Non-metals

Subtopics: Reaction between Metals and Non-metals

Questions

Q1. Write the electron dot structures for sodium, oxygen and magnesium.

Q2. Show the formation of Na_2O and MgO by the transfer of electrons. What are ions present in these compounds?

Q3. Why do ionic compounds have high melting points ?

Q4. Write the electron dot structures for potassium and chlorine.

Q5. Show the formation of KCl by the transfer of electrons.

Q6. (a) Show the formation of CaS by the transfer of electrons.

(b) Name the ions present in this compound CaS .

Q7. Explain the formation of ionic compound CaO with electron dot structure.

Q8. Why are ionic compounds usually hard?

Q9. How is it that ionic compounds in the solid state do not conduct electricity but they do so when in molten state?

Q10. Why are aqueous solutions of ionic compounds able to conduct electricity?